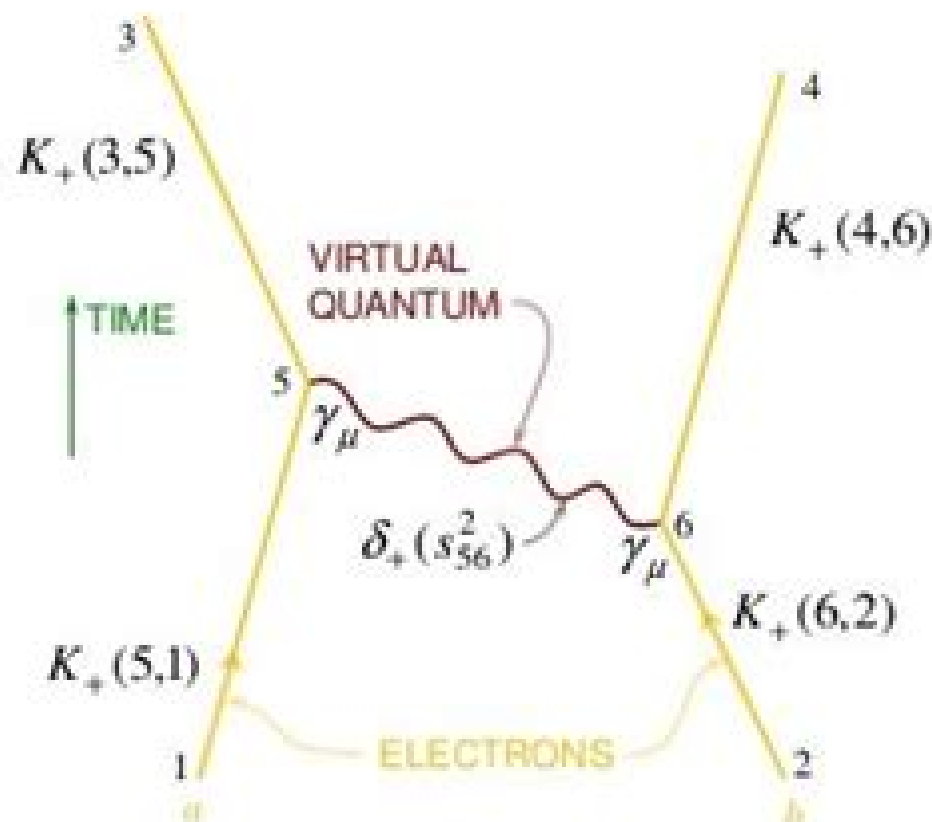


PART VII – QUANTUM ELECTRODYNAMICS



$$K'(3,4;1,2) = (-ie^2) \iint K_{+a}(3,5) K_{+b}(4,6) \gamma_{a\mu} \gamma_{b\mu} \delta_+(s_{56}^2) K_{+a}(5,1) K_{+b}(6,2) d\tau_5 d\tau_6$$

Circa 1949

Chapter 2 From First Principles

January 2017 – R4.4
Maurice R. TREMBLAY

Quantum Electrodynamics Part I

Puja Mehta



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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the

copper ii is then reduced and gains electrons

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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