

Air and Water Pollution Prevention and Control Engineering

Homework # 2

Answers to Selected Problems

1.- At the following link: <http://www.cwp.rpi.edu/~ernesto/F2013/AWPPCE/AdditionalReadings/Silva2013-ERL-APMortality.pdf> you will find a recent journal article reporting research on mortality associated with air pollution. Read the paper carefully through and create a ~ 250 word critical essay relating the contents of the paper to the notion of NAAQS.

Answer: Individual student answers.

2.- The 1-hour NAAQS for carbon monoxide is 40 mg/m³. Convert this to ppm at standard temperature y pressure.

Answer: $C = 40 \text{ mg/m}^3 = 0.04 \text{ g/m}^3$, dividing by $M_{CO} = 28 \text{ g/mol}$ yields $c = C/M = 0.00142 \text{ mol/m}^3$. From the ideal gas law, the volume of 1 mol of air is $V = RT/p = 8.3144 (298.15)/101,325 = 0.0244 \text{ m}^3$ and the number of mols in 1 m³ of air is $1/V = 40.87 \text{ mol/m}^3$. The concentration in ppm is then $(0.00142/40.87) \times 10^6 = 34.7 \text{ ppm}$.

3.- A solid waste incinerator emits particulate matter at the NSPS rate of 0.18 g/dscm. The incinerator burns 50 metric tons per day and exhausts gases at a ratio of 20 kg per kg of feed at atmospheric pressure and 453 K. Assume that the average molecular weight of the emitted gases is 30 and that they contain 12% CO₂ and 10% H₂O. What is the daily emission rate of particulate matter?

Answer: Note that dscm stands for "dry standard cubic meter". To obtain "standard cubic meter" of gas must multiply 0.18 by 0.9. To obtain the m³ at the not standard temperature of 453, must multiply by (298/453). This yields the particulate emission concentration as: $0.18 (0.9)(298/453) = 0.1065 \text{ g/m}^3$.

From the ideal gas law, the volume of 1 mol of exhaust gas is $V = RT/p = 8.3144 (453)/101,325 = 0.037 \text{ m}^3/\text{mol}$ and the volume per g is $0.037/30 = 0.00123 \text{ m}^3/\text{g}$ and the number of m³ of exhaust gas produced per day is obtained as $(0.00123)(50 \times 10^6)(20) = 1.230 \times 10^6 \text{ m}^3/\text{d}$. Therefore the amount of particulate matter emitted per day is $= 0.1065(1.230 \times 10^6) = 131,350 \text{ g/d}$.

4.- Methanol (CH₃OH) is burned in dry air at an equivalence ratio of 0.75.

(a) Determine the fuel-air mass ratio.

(b) Determine the composition of the combustion products.

Answer: The simplified combustion reaction of methanol with oxygen is

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L Darling-Hammond



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