



# Lost Isobel Ree

**J Dewey**



**Lost Isobel Ree:**

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web for example take the reaction of lead ii nitrate with hydrochloric acid to form lead ii chloride and nitric acid shown below  

$$\text{Pb(NO}_3)_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{HNO}_3(\text{aq})$$

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web answer key to practice problems on net ionic equations molecular  $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$  total ionic  $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$

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web you may have to consult the solubility rules write the net ionic equation for the reaction of  $\text{BaCl}_2(\text{aq})$  and  $\text{Na}_2\text{SO}_4(\text{aq})$  you may have to consult the solubility rules write

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precipitate will occur in a reaction to do so you can

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web choose the correct net ionic equation representing the reaction that occurs when solutions of potassium carbonate and copper i chloride are mixed a  $\text{K}_2\text{CO}_3 \text{ aq}$   $2\text{CuCl} \text{ aq}$

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web the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in

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web the net ionic equation for equation 1 is given as equation 3  $\text{s}_2 \text{ aq}$   $\text{Fe}_2 \text{ aq}$   $\text{gt}$   $\text{fes s}$  eq 3 8 note that the phase indicator aq is optional on ionic species but is

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web  $\text{CuClO}_4$  2 is soluble all perchlorates are soluble while  $\text{PbSO}_4$  is insoluble sulfates are generally soluble except for  $\text{Pb}^{2+}$   $\text{Ba}^{2+}$   $\text{Hg}_2^{2+}$  and  $\text{Sr}^{2+}$  the net ionic equation is

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web balanced chemical equation  $2\text{NaOH (aq)} + \text{MgCl}_2\text{ (aq)} \rightarrow 2\text{NaCl (aq)} + \text{Mg(OH)}_2\text{ (s)}$  full ionic equation  $2\text{Na}^+\text{ (aq)} + 2\text{OH}^-\text{ (aq)} + \text{Mg}^{2+}\text{ (aq)} + 2\text{Cl}^-\text{ (aq)} \rightarrow 2\text{Na}^+\text{ (aq)} + 2\text{Cl}^-\text{ (aq)} + \text{Mg(OH)}_2\text{ (s)}$

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