

R<sub>3</sub>CH 1.4 - 1.7

R−C=C−CH<sub>3</sub> 1.6 - 2.6

$$\stackrel{\triangleright}{\mathbb{R}}$$
 R−C−CH<sub>3</sub> 2.1 - 2.5

 $\stackrel{\triangleright}{\mathbb{R}}$  RO−C−CH<sub>3</sub> 2.1 - 2.6

N≡C−CH<sub>3</sub> 2.1 - 3.0

 $\stackrel{\triangleright}{\mathbb{R}}$  1.7 - 2.7

R−N−C−H 2.2 - 2.9

R−S−C−H 2.0 - 3.0

This is also true for any proton capable of making hydrogen bonding:

R—SH 1.0 - 5.0

R = alkyl or aryl

R—NH<sub>2</sub> 1.0 - 5.0

1°, 2°

R—OH 1.0 - 5.0

1°, 2°, 3°

—OH 4.0 - 7.0

11 - 12

Downfield shifts more common

very dry.

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**Lei Huang** 



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