

Trigonometry

Ratios and Laws

Trigonometric Ratios



$$\tan \alpha = \frac{a}{b}$$

$$\tan \beta = \frac{b}{a}$$

$$\cot \alpha = \frac{1}{\tan \alpha} = \frac{b}{a}$$

$$\cot \beta = \frac{1}{\tan \beta} = \frac{a}{b}$$

$$\tan \theta = \frac{\sin \theta}{\cos \theta}$$

$$\sin \alpha = \frac{a}{c}$$

$$\sin \beta = \frac{b}{c}$$

$$\csc \alpha = \frac{1}{\sin \alpha} = \frac{c}{a}$$

$$\csc \beta = \frac{1}{\sin \beta} = \frac{c}{b}$$

$$\cot \theta = \frac{\cos \theta}{\sin \theta}$$

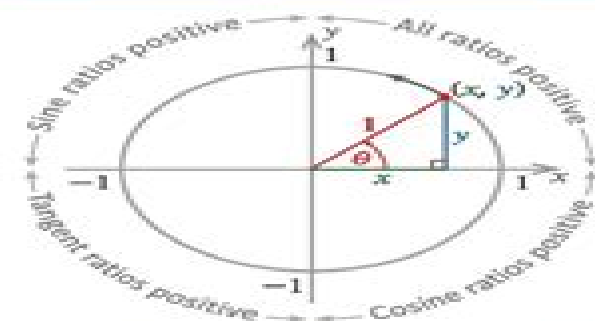
$$\cos \alpha = \frac{b}{c}$$

$$\cos \beta = \frac{a}{c}$$

$$\sec \alpha = \frac{1}{\cos \alpha} = \frac{c}{b}$$

$$\sec \beta = \frac{1}{\cos \beta} = \frac{c}{a}$$

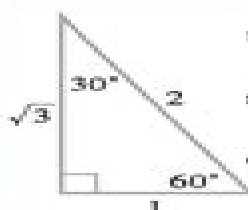
Unit Circle and Signs in Quadrants



$$\tan \theta = \frac{y}{x} \quad \sin \theta = \frac{y}{1} = y \quad \cos \theta = \frac{x}{1} = x$$

$$\cot \theta = \frac{x}{y} \quad \csc \theta = \frac{1}{y} \quad \sec \theta = \frac{1}{x}$$

Exact Ratios



$$\tan 60^\circ = \sqrt{3}$$

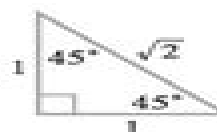
$$\tan 30^\circ = \frac{1}{\sqrt{3}}$$

$$\sin 60^\circ = \frac{\sqrt{3}}{2}$$

$$\sin 30^\circ = \frac{1}{2}$$

$$\cos 60^\circ = \frac{1}{2}$$

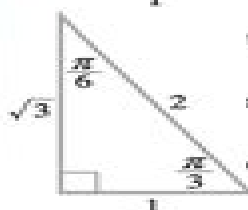
$$\cos 30^\circ = \frac{\sqrt{3}}{2}$$



$$\tan 45^\circ = 1$$

$$\sin 45^\circ = \frac{1}{\sqrt{2}}$$

$$\cos 45^\circ = \frac{1}{\sqrt{2}}$$



$$\tan \frac{\pi}{3} = \sqrt{3}$$

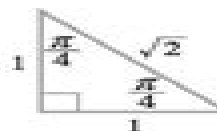
$$\tan \frac{\pi}{6} = \frac{1}{\sqrt{3}}$$

$$\sin \frac{\pi}{3} = \frac{\sqrt{3}}{2}$$

$$\sin \frac{\pi}{6} = \frac{1}{2}$$

$$\cos \frac{\pi}{3} = \frac{1}{2}$$

$$\cos \frac{\pi}{6} = \frac{\sqrt{3}}{2}$$

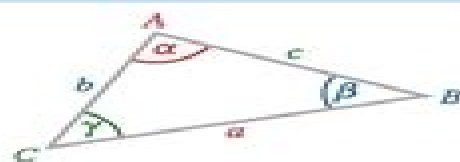


$$\tan \frac{\pi}{4} = 1$$

$$\sin \frac{\pi}{4} = \frac{1}{\sqrt{2}}$$

$$\cos \frac{\pi}{4} = \frac{1}{\sqrt{2}}$$

Sine Law for Triangles



$$\frac{a}{\sin \alpha} = \frac{b}{\sin \beta} = \frac{c}{\sin \gamma}$$

$$\text{Area} = \frac{1}{2} ab \sin \gamma = \frac{1}{2} ac \sin \beta = \frac{1}{2} bc \sin \alpha$$

$$\alpha = \sin^{-1} \left(\frac{a \sin \beta}{b} \right) = \sin^{-1} \left(\frac{a \sin \gamma}{c} \right)$$

$$\beta = \sin^{-1} \left(\frac{b \sin \alpha}{a} \right) = \sin^{-1} \left(\frac{b \sin \gamma}{c} \right)$$

$$\gamma = \sin^{-1} \left(\frac{c \sin \alpha}{a} \right) = \sin^{-1} \left(\frac{c \sin \beta}{b} \right)$$

Cosine Law for Triangles

$$a^2 = b^2 + c^2 - 2bc \cos \alpha \quad \alpha = \cos^{-1} \left(\frac{b^2 + c^2 - a^2}{2bc} \right)$$

$$c^2 = a^2 + b^2 - 2ab \cos \gamma \quad \gamma = \cos^{-1} \left(\frac{a^2 + b^2 - c^2}{2ab} \right)$$

$$b^2 = a^2 + c^2 - 2ac \cos \beta \quad \beta = \cos^{-1} \left(\frac{a^2 + c^2 - b^2}{2ac} \right)$$

Tangent Law for Triangles

$$\frac{a-b}{a+b} = \frac{\tan \left[\frac{1}{2} (\alpha - \beta) \right]}{\tan \left[\frac{1}{2} (\alpha + \beta) \right]}$$

$$\frac{b-c}{b+c} = \frac{\tan \left[\frac{1}{2} (\beta - \gamma) \right]}{\tan \left[\frac{1}{2} (\beta + \gamma) \right]}$$

$$\frac{a-c}{a+c} = \frac{\tan \left[\frac{1}{2} (\alpha - \gamma) \right]}{\tan \left[\frac{1}{2} (\alpha + \gamma) \right]}$$

Mollweide's Rule

$$\frac{a+b}{c} = \frac{\cos \left[\frac{\alpha - \beta}{2} \right]}{\sin \left[\frac{\gamma}{2} \right]}$$

$$\frac{a-b}{c} = \frac{\sin \left[\frac{\alpha - \beta}{2} \right]}{\cos \left[\frac{\gamma}{2} \right]}$$

Degrees and Radians

If θ is an angle in degrees and α is the same angle in radians:

$$\theta = \alpha \times \frac{180}{\pi}$$

$$\alpha = \theta \times \frac{\pi}{180}$$

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half reaction balance the atom undergoing redox changes if necessary add the number of electrons that correspond to the change in oxidation state

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b ba2 c br2 d ca e ga3 f h2 g h

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe} + \text{Co} \rightarrow \text{Co} + 2 \text{ Fe}$
 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu} + \text{Pb} \rightarrow \text{Pb} + 2 \text{ Cu}$ oxidation reduction 26 O_2

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 $5\text{Co} + \text{i}_2 \rightarrow 5\text{Co}^{2+} + \text{i}_2$ $\text{o}_5 + 5\text{C} \rightarrow \text{o}_2 + 5\text{C}^{2+}$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^{2+} oxidized reducing agent O_2 to O^{2-} reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^{+} oxidized red ag S^0 to S^{2-} reduced ox ag $3\text{C} + 2\text{Br}_2$

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