

1. From part B, clearly show how you calculated T_f^* .

$$-y(m_2 - m_1) = m_2b_1 - m_1b_2$$

$$-y(-0.47437 - (-2.7067)) = (-0.47437 \times 44.108) - (-2.7067 \times 38.949)$$

$$y = 37.8526 \text{ } ^\circ\text{C}$$

2. From part B, clearly Show how you calculated the value for freezing point depression $\Delta T_f = T_f - T_f^*$.

$$\Delta T_f = T_f - T_f^*$$

$$\Delta T_f = 42.15 - 37.85$$

$$\Delta T_f = 4.3 \text{ } ^\circ\text{C}$$

3. Calculate the molar mass of your unknown. Show the formula used, data, and units in your calculations.

$$\text{Molar mass of solute (g/mol)} = \frac{\text{mass of solute (g)} \times K_f \left(^\circ\text{C} \frac{\text{kg}}{\text{mol}} \right)}{\text{mass of solvent (kg)} \times \Delta T_f \left(^\circ\text{C} \right)}$$

$$\text{Molar mass of solute (g/mol)} = \frac{(0.404 \text{ g})(3.9 \text{ } ^\circ\text{C} \frac{\text{kg}}{\text{mol}})}{(2.882 \text{ g} \times \frac{1 \text{ kg}}{1000})(4.3 \text{ } ^\circ\text{C})}$$

$$\text{Molar mass of solute (g/mol)} = 176.045 = 176$$

4. Your unknown is either camphor or naphthalene. Identify your unknown by comparing its calculated molar mass to the molar masses of these two organic compounds. You will have to look up their chemical formulas and calculate the molar masses for each of the possible compounds.

Camphor molar mass: 152.23 g/mol

Naphthalene molar mass: 128.17 g/mol

Since the molar mass of my unknown was closer to 152.23 than 128.17 g/mol, the identity of my unknown was most likely camphor.

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