

Chapter 13: Colligative Properties of Solutions**Sections to read before coming to class:**

13-1, 13-2, 13-3, 13-4, 13-5, and 13-6

I. Objectives

By the end of this chapter, you should be able to describe:

- A. Definitions
- B. Concentration units
 - 1. Molar concentration and molal concentration
 - 2. Mole fraction and mole percent
 - 3. Weight fraction and weight percent
- C. The Solution Process
 - 1. Heats of solution
 - 2. "Like dissolves like"
 - 3. Factors affecting solubility
- D. Colligative properties
 - 1. Vapor pressure and Raoult's Law
 - 2. Boiling point elevation and freezing point depression
 - 3. Molar masses from colligative properties
 - 4. Osmotic pressure

II. Definitions

- A. **Solution:** homogenous mixture of a solute dissolved in a solvent (or a solvent mixture).

- B. **Saturated solution:** Dissolved solute is in equilibrium with solid solute $\text{solute}_{(aq)} \rightleftharpoons \text{solid}$

- C. **Solubility:** Amount of a solute req'd to make a saturated solution (temperature dependent)

- D. **Unsaturated solution:** Solution has less solute than required to make a saturated solution.

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