

CHM 312: INORGANIC CHEMISTRY II

LECTURE NOTES

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CHEMISTRY OF FIRST ROW TRANSITION METALS

The first row transition metals consist of elements: Scandium (Sc), Titanium (Ti), Vanadium (V), Chromium (Cr), Manganese (Mn), Iron (Fe), Cobalt (Co), Nickel (Ni), Copper (Cu), and Zinc (Zn) as in the periodic table. These elements have partially filled d orbitals and exhibit a wide range of physical and chemical properties, making them important in many areas of chemistry, including catalysis, materials science, and bioinorganic chemistry. The partially filled d orbitals gives them unique properties such as high melting and boiling points, good electrical conductivity, and the ability to form complex ions and compounds. These properties arise from the ability of the d orbitals to participate in bonding, allowing the metals to form a variety of coordination complexes.

Electronic Configurations:

The electronic structure of the first row transition elements is defined by the filling of the 3d and 4s orbitals. These elements are characterized by their partially filled d-orbitals, which play a crucial role in their chemical and physical properties. These elements have electronic configurations that follow the general pattern $[\text{Ar}] 3d^{1-10} 4s^{1-2}$. The electronic configuration of the first row transition elements using the s, p, d, and f methods:

Scandium (Sc):

s method: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

p method: $[\text{Ar}] 3d^1 4s^2$

d method: $[\text{Ar}] 3d^1 4s^2$

f method: $[\text{Ar}] 3d^1 4s^2 4p^0 4d^0 4f^0$

Titanium (Ti):

s method: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

p method: $[\text{Ar}] 3d^2 4s^2$

d method: $[\text{Ar}] 3d^2 4s^2$

f method: $[\text{Ar}] 3d^2 4s^2 4p^0 4d^0 4f^0$

Vanadium (V):

s method: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$

p method: $[\text{Ar}] 3d^3 4s^2$

d method: $[\text{Ar}] 3d^3 4s^2$

f method: $[\text{Ar}] 3d^3 4s^2 4p^0 4d^0 4f^0$

Chromium (Cr):

s method: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$

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