

## IBFT 2018 Methodical Practice/Packet for Chapter 7 and 8 (PLASMA)

### Multiple Choice

Simple formulae for the complex formulae is shown in parenthesis.

- \_\_\_\_\_ 1. The transmittance of a pure  $\text{Ca}^{2+}$  solution is 80% relative to the pure solvent. The optical density of a pure  $\text{Ca}^{2+}$  solution will be by 95% compared to the solvent  $\text{Ca}^{2+}$  solution by?
  - a. 14%
  - b. 20%
  - c. 65%
  - d. 88%
  
- \_\_\_\_\_ 2. The equilibrium of the reaction  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  is  $K_w = 1.0 \times 10^{-14}$  at 25°C. The concentration of  $\text{H}_3\text{O}^+$  in a 0.1 M  $\text{H}_2\text{O}$  solution is?
  - a.  $1.0 \times 10^{-14}$
  - b.  $1.0 \times 10^{-13}$
  - c.  $1.0 \times 10^{-12}$
  - d.  $1.0 \times 10^{-11}$
  
- \_\_\_\_\_ 3. A solution of  $\text{H}_2\text{O}$  is prepared by mixing 100 mL of 0.1 M  $\text{H}_2\text{O}$  with 100 mL of 0.1 M  $\text{H}_2\text{O}$ . The pH of the solution is?
  - a.  $\text{pH} = 1.0$  at 25°C
  - b.  $\text{pH} = 1.0$  at 25°C
  - c.  $\text{pH} = 1.0$  at 25°C
  - d.  $\text{pH} = 1.0$  at 25°C
  
- \_\_\_\_\_ 4. A solution of  $\text{H}_2\text{O}$  is prepared by mixing 100 mL of 0.1 M  $\text{H}_2\text{O}$  with 100 mL of 0.1 M  $\text{H}_2\text{O}$ . The pH of the solution is?
  - a.  $\text{pH} = 1.0$  at 25°C
  - b.  $\text{pH} = 1.0$  at 25°C
  - c.  $\text{pH} = 1.0$  at 25°C
  - d.  $\text{pH} = 1.0$  at 25°C
  
- \_\_\_\_\_ 5. A solution of  $\text{H}_2\text{O}$  is prepared by mixing 100 mL of 0.1 M  $\text{H}_2\text{O}$  with 100 mL of 0.1 M  $\text{H}_2\text{O}$ . The pH of the solution is?
  - a.  $\text{pH} = 1.0$  at 25°C
  - b.  $\text{pH} = 1.0$  at 25°C
  - c.  $\text{pH} = 1.0$  at 25°C
  - d.  $\text{pH} = 1.0$  at 25°C

# Rspt Verbal Test 3 Pk1

**D Siedentop**



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