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**Camilla Rothe**



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determine based on electronegativity and symmetry if the following molecules are polar or nonpolar hcn h c n cf4 c f 4 cocl2  
c o c l 2 nf3 n f 3 q3

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web the relative electronegativities of the bonded atoms is  $\text{H} < \text{C} < \text{Cl}$  and so the bond moments all point toward the  $\text{Cl}$  end of the molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as  $\text{BF}_3$  trigonal planar  $\text{CH}_4$  tetrahedral  $\text{PF}_5$  trigonal bipyramidal and  $\text{SF}_6$

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