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web log [H⁺] = 10^{-4.05} = 8.91 × 10⁻⁵ mol dm⁻³ log pK_a = pH = 3.75 + 4.05 = 7.80 and HCOOH = 0.1 so HCOO⁻ = 0.200 mol dm⁻³ amount vol x

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
















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