

# MACHINES



THE HANDS-ON APPROACH TO SCIENCE

# Machines Make It Work Science Series

**J Elliott**



## **Machines Make It Work Science Series:**

## Decoding **Machines Make It Work Science Series**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Machines Make It Work Science Series**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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web 1 consider the reaction  $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$  a 80.0 grams of iodine(V) oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced b if in the above situation only 0.160 moles of iodine  $\text{I}_2$  was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C<sub>3</sub>H<sub>8</sub> what is the percent yield  
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation  $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$   
 8.3 mole H<sub>2</sub>S  $\times$  18g H<sub>2</sub>O / 1 mole H<sub>2</sub>O = 149.4g H<sub>2</sub>O is your theoretical yield so percent yield =  $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$   
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH<sub>3</sub> can be produced theoretically from the reaction of 5.0 g of N<sub>2</sub> what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$  if the reaction of 21.3 grams of O<sub>2</sub> produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  $\frac{21.3 \text{ g O}_2}{32 \text{ g/mol}} = 0.666 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design  $\text{yield} = \frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> from 160.0 g Al  $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.93 \text{ mol Al}$   $\frac{4 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} = 1.48 \text{ mol Al}_2\text{O}_3$   $1.48 \text{ mol Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 150.9 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{126.0 \text{ g}}{150.9 \text{ g}} \times 100 = 83.5\%$

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield =  $\frac{86.6 \text{ g}}{100 \text{ g}} \times 100 = 86.6\%$  only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H<sub>2</sub> are left over 13

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