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# Police Clerk Passbooks For Career Opportunities

**J Elliott**



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mass  $\text{C}_6\text{H}_6$  6 12 6 1 g mol 1 78 g mol 1 hence  $\text{C}_6\text{H}_6$  s number of moles  $\frac{30.78 \text{ g}}{78 \text{ g mol}^{-1}} = 0.3946 \text{ mol}$  molar mass of carbon tet  $\text{CCl}_4$  12 4 355 154 g mol 1

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web solution moles of  $\text{CH}_3\text{OH}$  5 5 32 0 17 mole moles of  $\text{H}_2\text{O}$  40 18 2 2 moles therefore according to the equation mole fraction of  $\text{CH}_3\text{OH}$   $\frac{0.17}{0.17 + 0.17} = 0.073$  to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute  $\frac{0.449 \text{ mol}}{2.92 \text{ kg}} \times 131.108 \text{ g mol}^{-1} = 15.6 \text{ g kbr}$

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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