Defining:

$$\epsilon_{\rm D} = \frac{\rm p(HDO)}{\rm p(H_2O)} - 1$$
; $\epsilon_{\rm O} = \frac{\rm p(H_2O^{18})}{\rm p(H_2O^{16})} - 1$

i. e., the fractional vapor pressure differences at a given temperature, and

$$E = \frac{R_2}{R_1} - 1$$

where E is, by standard terminology, the total phase enrichment in either heavy isotope, R_1 is the initial molecular ratio HDO/H_2O or H_2O^{18}/H_2O^{16} in the phase, and R_2 is the final ratio, then for the enrichments of deuterium and O^{18} in the liquid phase one readily derives

$$E_D = f^{\epsilon_D} - 1$$
 $E_O = f^{\epsilon_O} - 1$ $\frac{E_D}{E_O} \sim \frac{\epsilon_D}{\epsilon_O}$

where f is the fraction of the liquid remaining. The enrichment ratio given by the last equation is a function of temperature alone and the approximation holds quite well for the small enrichments encountered in natural systems.

Now since H₂O¹⁶ is the most volatile species below temperatures of several hundred degrees, the liquid phase will become progressively enriched in the heavy isotopes during an equilibrium evaporation process, or the vapor phase will become progressively depleted in the heavy isotopes during an equilibrium condensation process, and according to the above equation the percentage enrichments or depletions in deuterium and Ol8 are related by the ratio of the fractional vapor pressure differences, the latter being known from experimental data. At room temperature the ratio is about 9, decreasing to about 5 at the boiling point of water. It appears, therefore, that if the meteoric water cycle of evaporation-condensation takes place under equilibrium conditions, the ocean waters should have about the highest concentrations of the heavy isotopes, while the terrestrial fresh waters should be relatively depleted in these isotopes, the extent of the depletion depending upon the fraction of the original water vapor remaining in the air mass from which their water has, on the average, been de-The depletion ratio, taking the depletions relative to average ocean water, should be of the order of 9, independent of the extent of depletion of either isotope.

This relationship was first established experimentally by the work of Epstein and Mayeda² and Friedman,³ who analyzed 36 marine and 9 terrestrial waters for both O¹⁸ and deuterium concentrations, and found a depletion ratio in the fresh waters corresponding to ordinary temperatures. Epstein and Mayeda further showed that the salinity-O¹⁸ variations in marine waters were consistent with a Rayleigh distillation process in which the oceanic precipitation is progres-

Nuclear Processes In Geologic Settings

Julia Schneider

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□ (Japan),197? **Progress in Isotope Geology** Kalervo Rankama,1963 Nuclear Processes In Geologic Settings

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