

Avogadro's Law:

under the same temperature and pressure, same volume of gases will consist of the same number of moles of gases.

Volume ratio = Mole ratio
of gases of gases

E.g. if 3 moles of gas A occupies 12 dm³, then 3 moles of gas B at the same temperature and pressure will also occupy 12 dm³.

number
of particles

$$\begin{array}{c} + \\ 6.02 \times 10^{23} \end{array} \quad \begin{array}{c} \times \\ 6.02 \times 10^{23} \end{array}$$

% PURITY =

$$\frac{\text{mass of PURE}}{\text{mass of IMPURE}} \times 100\%$$

% YIELD =

$$\frac{\text{EXPERIMENTAL mass}}{\text{CALCULATED mass}} \times 100\%$$

Limiting reagent: check when no. of moles of all reactants are given

Qn: 0.200 mol of H₂ was reacted with 0.400 mol of O₂ to produce water. Determine the limiting reagent of the reaction.

Ans: 0.200 mol of H₂ required 0.100 mol of O₂ for complete reaction. However 0.400 mol of O₂ was supplied. Therefore O₂ is in excess and H₂ is the limiting reagent.

**volume (dm³)
of gas at r.t.p.**

$$\begin{array}{c} + 24 \text{ dm}^3 \\ \times 24 \text{ dm}^3 \end{array}$$

number of
MOLES

$$\begin{array}{c} \times M_r \\ + M_r \end{array}$$

**mass (g)
of substance**

Dilution:

number of moles stays the same,
total volume of solution increases,
concentration of solution decreases

$$\text{conc. (diluted)} = \text{conc. (original)} = \frac{\text{vol. (original)}}{\text{vol. (new)}}$$

$$\begin{array}{c} + \text{vol (dm}^3\text{)} \\ \times \text{vol (dm}^3\text{)} \end{array} \text{ of solution}$$

$$\begin{array}{c} + \\ \times \end{array} \text{vol (dm}^3\text{)} \text{ of solution}$$

Sampling:

number of moles decreases,
volume of solution decreases,
concentration remains the same

**conc. of solution
(mol dm⁻³)**

$$\begin{array}{c} \times M_r \\ + M_r \end{array}$$

**conc. of solution
(g dm⁻³)**

Mole Concept In Chemistry

Singhal



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