

d. A natural resource for humans and other organisms

Water is crucial for life on Earth. It is an essential reactant in the production of glucose through photosynthesis as shown: $6\text{CO}_2(g) + 6\text{H}_2\text{O}(l) + \text{light} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(aq) + 6\text{O}_2(g)$

Water also plays a role in cellular respiration, providing energy for life:



Water also forms a large part of the human body, being over 80% water. Without water, blood would not be able to flow and transport nutrients and waste products. Water also facilitates the removal of waste products from the body.

Water is also used by humans for drinking, washing, agriculture and industrial processes.

5. Construct Lewis electron dot structures of water, ammonia and hydrogen sulfide to identify the distribution of electrons



6. Compare the molecular structure of water, ammonia and hydrogen sulfide, the differences in their molecular shapes and in their melting and boiling points

MOLECULE	MOLECULAR SHAPE	MELTING POINT (°C)	BOILING POINT (°C)
H_2O	Bent	0	100
NH_3	Trigonal Pyramidal	-78	-33
H_2S	Bent	-86	-60

7. Describe hydrogen bonding between molecules

Hydrogen bonding is a special form of dipole-dipole attraction. Dipole-dipole forces refer to the force of attraction between the oppositely charged ends of polar molecules. This means that if two polar molecules approach, they will tend to orient themselves such that the positive end of one molecule is close to the negative end of another. The ends of a dipole have small charge, hence dipole-dipole forces are comparatively weak.

Hydrogen bonding occurs when a hydrogen atom is bonded to an atom of oxygen, nitrogen or fluorine. As these elements are highly electronegative, the shared electrons tend to spend more time around these atoms, leaving hydrogen electron-less and positively charged. Hence, this hydrogen is attracted to the negative oxygen, nitrogen and fluorine atoms of other molecules, forming a hydrogen bond.

Hydrogen bonds are about 10 times stronger than dipole-dipole bonds but 10 times weaker than ionic or covalent bonds.

Modules Of Chem Struct Dynprelim

Patrick Vollmar



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