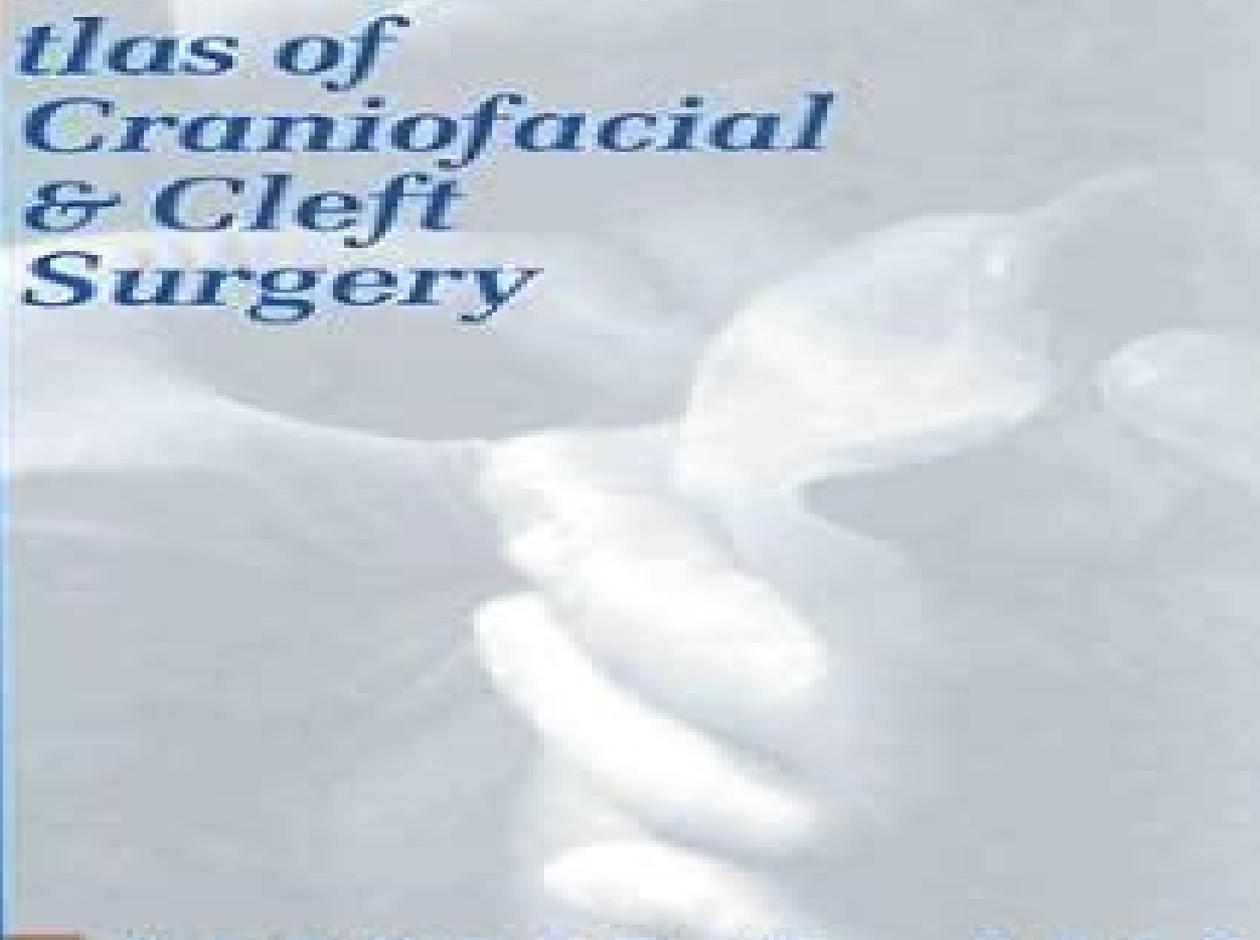


SALYER AND BARDACH



*Atlas of
Craniofacial
& Cleft
Surgery*

VOLUME I
CRANIOFACIAL
SURGERY

Missouri - Elsevier

KENNETH E. SALYER

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 $8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8 3 \text{ mole } \text{H}_2\text{O}$
149 4g H_2O is your theoretical yield so percent yield $\frac{137 1\text{g } \text{H}_2\text{O}}{149 4\text{g } \text{H}_2\text{O}} \times 100 = 91 77\%$
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 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$
if the reaction of 21 3 grams of O_2 produces a 47 3 yield how many grams of CO would be produced
step one figure out how many moles of the limiting reagent you have
 $\frac{21 3 \text{ grams } \text{O}_2}{32 \text{ g/mole}} = 0 66 \text{ moles } \text{O}_2$

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 $\frac{\text{calculated mass of the product}}{\text{mass of the product expected}} \times 100$
in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
b determine the theoretical yield of Al_2O_3 160 0

g al x 1 mol al x 2 mol al₂O₃ x 101 964 g al₂O₃ 302 3 g al₂O₃ 26 982 g al 4 mol al 1 mol al₂O₃ c determine the percent yield
yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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c is the limiting reagent 4 33 g of h 2 are left over 13

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