

SALYER AND BARDACH



Atlas of Craniofacial & Cleft Surgery

VOLUME I
CRANIOFACIAL
SURGERY

Mosby - Elsevier

KENNETH E. SALLY

Salyer And Bardachs Atlas Of Craniofacial And Cleft Surgery Volume 1 Craniofacial Surgery

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation
$$8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$$

149 4g H_2O is your theoretical yield so percent yield
$$\frac{137 1 \text{ g } \text{H}_2\text{O}}{149 4 \text{ g } \text{H}_2\text{O}} \times 100 = 91 77\%$$

91 77 problems 6 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21 3 grams of O_2 produces a 47 3 yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
$$\frac{21 3 \text{ g } \text{O}_2}{32 \text{ g } \text{O}_2} = 0 66 \text{ moles } \text{O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained
$$\frac{\text{calculated mass of the product}}{\text{mass of the product expected}} \times 100$$

in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction
$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al_2O_3 160 0

$\frac{\text{g Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3}{302.3 \text{ g Al}_2\text{O}_3 \times 26.982 \text{ g Al} \times 4 \text{ mol Al} \times 1 \text{ mol Al}_2\text{O}_3}$ determine the percent yield
yield actual $\times 100$
 $\frac{260.0 \text{ g}}{100.8601 \text{ theoretical}} \times 302.3 \text{ g}$

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80 g I_2O_5 1 mol I_2O_5 1 mol I_2 x 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

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