

Периодический закон в свете теории строения атома

Периодический закон занимает центральное место в химии, являясь фундаментальным принципом, который раскрывает взаимосвязь между атомным строением элементов и их химическими свойствами. Понимание периодического закона в контексте теории строения атома позволяет глубже осознать причины периодичности физических и химических характеристик элементов, а также прогнозировать и объяснять многие явления в химии.

Краевская Дарья 11 Б

Теория строения атома

Модель атома Резерфорда

Согласно модели Резерфорда, атом состоит из положительно заряженного ядра, вокруг которого вращаются отрицательно заряженные электроны. Эта модель помогла объяснить многие наблюдаемые явления, но не могла дать полное представление о внутренней структуре атома.

Модель атома Бора

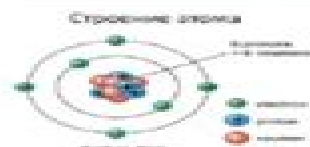
Модель Бора постулировала, что электроны в атоме могут находиться только на определенных стационарных энергетических уровнях. Переходя с одного уровня на другой, электрон поглощает или испускает фотон определенной энергии, что объясняет дискретность спектра атома.

Квантово-механическая модель

Современная квантово-механическая модель атома описывает электрон как волну, вероятность нахождения которой в пространстве вокруг ядра характеризуется квантовыми числами. Это позволяет объяснить многие особенности строения и свойств атомов.



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