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Rachel S Tattersall



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web apr 7 2014 molarity moles of solute liters of solution 8.422 a first convert 250 ml to liters 250/1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M 3. c a solution with molarity 2 requires 2 M of NaOH per liter so 4 x 2.8 M = 11.2 M a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture M_{mix} can be calculated using the following formula $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$ $\frac{480 \text{ g} \times 1.2 \text{ L} + 520 \text{ g} \times 2.0 \text{ L} + 624 \text{ g} \times 1.0 \text{ L} + 1000 \text{ g} \times 1.344 \text{ L}}{1.2 \text{ L} + 2.0 \text{ L} + 1.0 \text{ L} + 1.344 \text{ L}} = 480 \text{ g}$ with increase in temperature which of these changes a molality b fraction of solute present in water

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