

# Radioisotopes in Hydrology

➤ A number of radioisotopes are used in hydrological studies.

➤ Determining the age of ground water is important for several Reasons -

- Groundwater age translates roughly to circulation and renewability, which is critical for water resource planning.
- The age of groundwater relates also to its vulnerability to contamination.
- Modern groundwaters will have more direct pathways from the surface, facilitating infiltration of contaminants from surface activities in the recharge area. Old groundwaters are isolated from the surface and perhaps less vulnerable to anthropogenic contamination.

TABLE 8.1

Tracers for Age Dating Groundwaters

Isotope	Half Life	Typical Half Time	Source (Parent Nuclide)	Typical Concentration
<b>Modern Groundwaters</b>				
$^3\text{H}$	12.32 y	<1 y	Cosmogenic ( $^3\text{H}_2$ )	Steady state
Tracers, $\text{F}$	41.02 y	<1 y	Thermoluminescent phosphors	1000-10 <sup>5</sup> or 10 <sup>6</sup> ppm
$^3\text{He}$	12.32 y		Cosmogenic ( $^3\text{H}$ )	0.01-10 <sup>3</sup>
$^3\text{H}_2$	12.32 y		Unknown, but evaporating	1.0-1000, increasing
$\text{SF}_6$	Stable	100 y	Electrical insulator emissions	2 ppm, increasing
CFC	Stable	100-200 y	Refrigerant emissions, aerosols	CFC-12: 100 ppb to 2000
<b>Sub Modern Groundwaters</b>				
$^3\text{H}_2$	120 y		Cosmogenic ( $^3\text{H}$ )	10-700 ppb
$^3\text{He}$	>12 y	>1 y	Cosmogenic ( $^3\text{H}$ )	2 ppm to 10 <sup>3</sup>
<b>Old Groundwaters/Minerals</b>				
$^{14}\text{C}$	5730 y	<1 y	Cosmogenic ( $^{14}\text{N}$ )	100-1000, 0.1-100 ppb/y <sup>2</sup>
$^3\text{H}_2$	120,000 y		Cosmogenic ( $^3\text{H}$ )	1-2 ppb/y <sup>2</sup>
$^{13}\text{C}$	500,000 y	<1 y	Cosmogenic ( $^{13}\text{N}$ )	<100 ppm to 10 <sup>3</sup>
			Subsurface ( $^{13}\text{C}$ )	$^{13}\text{C}/^{12}\text{C} = 989.8 \times 10^{-6}$
$^{18}\text{O}$	11.4 My	<1 y	Cosmogenic ( $^{18}\text{N}$ )	100,000 ppm to 10 <sup>6</sup>
			Subsurface ( $^{18}\text{O}$ ) up to 10000	>1000 ppm to 10 <sup>4</sup>
$\text{H}_2$	Stable		Subsurface ( $\text{H}_2$ ) and $\text{H}_2$ in decay	1.20x10 <sup>3</sup> ppm
$^3\text{H}_2$	Stable		Subsurface ( $^3\text{H}_2$ )	0.01%
$^3\text{He}$	Stable		Subsurface ( $^3\text{H}$ ) up to 1000	0.01% of $^3\text{He}$

Clark, 2015

# Radioisotopes In Hydrology

**M Planty**



## **Radioisotopes In Hydrology:**

Radioisotopes in Hydrology International Atomic Energy Agency,1963      **Practical Isotope Hydrology** S. M. Rao,2006  
In Indian context      Isotope Techniques in Hydrology International Atomic Energy Agency,1968      **Introduction to Isotope Hydrology** Willem Gerrit Mook,2006 This publication presents in a simple but scientifically rigorous manner the consequences of natural processes to stable isotope concentrations and radioactivities and discusses the implications of these processes      *Isotopes in Hydrology, Hydrogeology and Water Resources* B. P. Singh,Bhishm Kumar,2005 Table of Contents  
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Agency,1963      *Isotope Hydrology* Pradeep Kumar Aggarwal,2012 Isotope Hydrology uses stable and radioactive isotopes of  
water and its dissolved constituents to trace hydrological processes including the pathways of rainfall and snowmelt to and  
interactions between aquifers lakes and rivers The potential of using stable isotopes of water was recognized in the 1930s  
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resource for graduate and post graduate level courses in hydrology reproducing many important papers which are otherwise  
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groundwater      *Guide to the Safe Handling of Radioisotopes in Hydrology* International Atomic Energy Agency,1966  
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March 1963. A Programme.* Symposium on the Application of Radioisotopes in Hydrology (INTERNATIONAL ATOMIC  
ENERGY AGENCY),1962      Isotope Hydrology: A Study Of The Water Cycle Joel R Gat,2010-01-08 Within the realm of the

newly evolving discipline of environmental sciences the stable isotope methodology is being used to an ever increasing extent especially in the study of the water cycle and of paleo climatology This book introduces the rules of the game by reviewing the natural variability of stable isotopes in the hydrosphere describing the physico chemical basis of isotope fractionation and applying this knowledge to natural waters as they move through the hydrologic cycle from the ocean to the atmosphere the biosphere and the lithosphere There is a special focus on the processes at the surface atmosphere and land biosphere atmosphere interfaces since these are the sites of major changes in isotope composition In response to the increasing awareness of our changing climate a discussion on the global view of the changing water cycle in the past and future winds up the presentation a

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