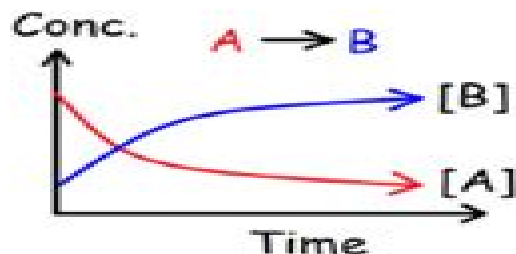


Chemical Kinetics:



Instantaneous Rate of Appearance:

$$\text{Rate} = + \frac{d[B]}{dt}$$

Method of Initial Rates:

Trial	[A]	[B]	[C]	I. Rate
1	0.10 M	0.10 M	0.10 M	0.20 M/s
2	0.20 M	0.10 M	0.10 M	0.40 M/s
3	0.10 M	0.20 M	0.10 M	0.80 M/s
4	0.10 M	0.10 M	0.20 M	0.20 M/s

Rate Constant k:

$$k = \frac{\text{Rate}}{[A]^x[B]^y[C]^z}$$

Units of k: $M^{1-n} t^{-1}$ or $(mol)^{1-n} (L)^{n-1} t^{-1}$

Note: $M = mol \cdot L^{-1}$ and $t \rightarrow s, min, hr, days$

Unimolecular:



Bimolecular:



Termolecular:



Average Rate of Appearance: $A \rightarrow B$

$$\text{Rate} = + \frac{\Delta[B]}{\Delta t} = \frac{[B]_f - [B]_i}{t_f - t_i}$$

Average Rate of Disappearance:

$$\text{Rate} = - \frac{\Delta[A]}{\Delta t}$$

Rate of a Chemical Reaction:



$$\text{Rate} = - \frac{1}{2} \frac{\Delta[A]}{\Delta t} = - \frac{1}{3} \frac{\Delta[B]}{\Delta t} = + \frac{1}{4} \frac{\Delta[C]}{\Delta t} = + \frac{1}{5} \frac{\Delta[D]}{\Delta t}$$

Differential Rate Law Expression:

$$\text{Rate} = k[A]^x[B]^y[C]^z$$

Finding The Order of a Reactant:

$$x = \frac{\log \left(\frac{\text{Rate 2}}{\text{Rate 1}} \right)}{\log \left(\frac{[A]_2}{[A]_1} \right)} \quad y = \frac{\log \left(\frac{\text{Rate 3}}{\text{Rate 1}} \right)}{\log \left(\frac{[B]_3}{[B]_1} \right)}$$

Overall order of the reaction:

$$\text{Order} = x + y + z$$

Factors Affecting the Rate of a Reaction:

1. Temperature
2. Concentration
3. Catalyst
4. Surface Area
5. The Nature of the Reactants

Reaction Mechanisms:



Note: $\text{Catalyst} \rightarrow B$ $\text{Intermediate} \rightarrow D$

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Terry C. Jones



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