

MATTER AND ITS CHANGES

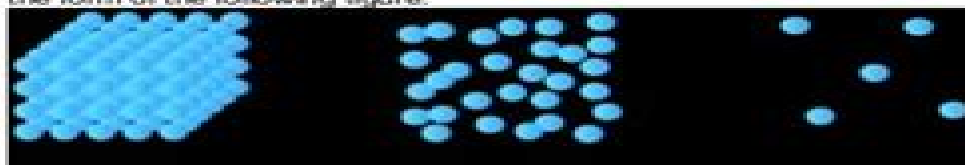
•EFFECT OF CHANGE IN TEMPERATURE ON THE STATES OF MATTER

We know that matter exists in three different physical states – solid, liquid, and gas. The states of matter are interchangeable i.e., a substance can be changed from its solid state to its liquid state and vice-versa. Transitions from the liquid state to the gaseous state and vice-versa are also possible. However, some substances directly change into their gaseous states from their solid states and vice-versa.

Water exists in all three states (i.e., solid, liquid, and gaseous) as ice, water, and water vapour respectively. These states are interchangeable (see the given figure).



Temperature is a major factor responsible for causing a change in the various states of matter. The changes in the states of matter with increasing or decreasing temperatures can be shown in the form of the following figure.



Effect of temperature on the states of matter

On heating a solid substance, the kinetic energy of its constituent particles increases. As a result, the particles start vibrating with greater speed. This extra energy helps the particles to overcome the inter-particle force of attraction. Soon, they leave their positions and start moving more freely. Consequently, the substance melts. The change of solid state into liquid state is known as fusion. The minimum temperature at which a substance starts melting and becomes a liquid at atmospheric pressure is known as its melting point.

It has been observed that temperature remains constant till all solid melts into liquid. The heat supplied is used for changing a solid state into its liquid state by overcoming the particle-particle attraction force. The solid substance absorbs heat energy without showing any rise in temperature. This heat energy is called latent heat. The amount of heat required to convert a unit mass of a solid into its liquid state, without a change in temperature (i.e., at its melting point), is called the latent heat of fusion.

For ice, the latent heat of fusion is 334 kJ/kg . This means that 334 kJ of heat is required to convert 1 kg of ice at 0°C into 1 kg of water at 0°C . Conversely, 334 kJ of heat is released when 1 kg of water freezes at 0°C to give 1 kg of ice at 0°C .

If heating is continued, then the kinetic energy of liquid particles increases further. This increases the velocity of the particles, and at a certain temperature, they obtain enough energy to break free from the force of attraction of the other particles. At this point, the liquid starts boiling i.e., starts

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