Ozone and Other Photochemical Oxidants

- Photochemical oxidants result from atmospheric interactions of nitrogen dioxide and sunlight
 - Most common is ozone O₃
 - Colorless gas with slightly sweet odor
 - Very active chemically, oxidizes or burns
 - Beneficial in the upper atmosphere

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United States. Environmental Criteria and Assessment Office

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Guderian,2012-12-06 Photochemical oxidants are secondary air pollutants formed under the influence of sunlight by complex photochemical reactions in air which contains nitrogen oxides and reactive hydrocarbons as precursors The most adverse components formed by photochemical reactions in polluted air are ozone 0 3 and peroxyacetyl nitrate PAN among many other products such as aldehydes ketones organic and inorganic acids nitrates sulfates etc An analysis and evaluation of the available knowledge has been used to characterize the relationships among emissions ambient air concentrations and effects and to identify the important controlling influences on the formation and effects of photochemical oxidants The biological activity of photochemical oxidants was first clearly manifested during the early 1940 s when vegetation injury was observed in the Los Angeles Basin in the United States Since that time as a consequence of the increasing emissions of photochemical oxidant precursors the photochemical oxidants have become the most important air pollutants in North America In other parts of the world for example South and Central America Asia and Australia photo chemical oxidants threaten vegetation particularly the economic and ecological performance of plant life According to my knowledge the first observations of ozone and PAN injury to vegetation in Europe were made by Dr Ellis F Darley Statewide Air Pollution Research Center University of California Riverside California during a study visit 1963 64 to the Federal Republic of Germany

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