

Principles of Macroeconomics

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web feb 14 2014 we have therefore started our studies by identifying a theory level that properly describes the competition between the pathways a d presented in fig 1 fig 1 possible reactions of toluene and with permanganate anion at positions a c m b c 1 c 2 c c 2 c 3 d c 3 c 4 full size image

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web aug 12 2022 kmno4 is a very powerful oxidant if you run it a high temperature with concentrated acid then phenylacetic acid loses co2 to give toluene which is then oxidised to benzoic acid same with styrene under controlled conditions you can get the diol push harder and that oxidises further waylander aug 12 2022 at 6 45

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web the oxidation of toluene by permanganate has been studied as a model for the oxidation of c h bonds by metal reagents metalloenzymes and metal oxide surfaces in water the reaction proceeds by hydride h transfer from toluene to a permanganate oxygen whereas in toluene solution permanganate abstracts a hydrogen atom h

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web may 15 2020 toluene 1 oxidation with KMnO_4 mechanism 2 benzoic acid $\text{C}_6\text{H}_5\text{COOH}$ SOCl_2 3 benzoyl chloride $\text{C}_6\text{H}_5\text{COCl}$ LiAlH_4 $\text{CH}_3\text{CH}_2\text{CH}_2\text{CHO}$ $\text{C}_6\text{H}_5\text{CHO}$ $\text{CH}_2\text{OHCH}_2\text{OH}$ h step 1 toluene oxidation with KMnO_4 KMnO_4 is a strong oxidising agent which oxidise toluene to benzaldehyde in 1st step

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web may 1 2004 oxidation of toluene using aqueous potassium permanganate was studied under heterogeneous condition in the presence of hydrodynamic cavitation and compared with the results of the reaction under acoustic cavitation

can acidified or neutral KMnO_4 oxidise toluene to benzoic acid - Jan 08 2023

web oct 15 2014 1 answer sorted by 6 here are the three equations describing the reduction of manganese and concurrent oxidation of whatever substrate may be present under basic neutral and acidic conditions respectively
 $\text{Mn}^{7+} + 4\text{e}^- + 4\text{H}^+ \rightarrow \text{Mn}^{3+} + 2\text{H}_2\text{O}$
 $\text{Mn}^{7+} + 3\text{e}^- + 4\text{H}^+ \rightarrow \text{Mn}^{4+} + 2\text{H}_2\text{O}$
 $\text{Mn}^{7+} + \text{e}^- + 4\text{H}^+ \rightarrow \text{Mn}^{6+} + 2\text{H}_2\text{O}$
basic $\text{Mn}^{7+} + 3\text{e}^- + 4\text{OH}^- \rightarrow \text{Mn}^{4+} + 2\text{H}_2\text{O}$
acidic $\text{Mn}^{7+} + 5\text{e}^- + 8\text{H}^+ \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}$

what is the mechanism of toluene oxidation by KMnO_4 reddit - Jul 14 2023

web mar 15 2019 what is the mechanism of toluene oxidation by KMnO_4 it is quite simple reaction that KMnO_4 can oxidize benzylic hydrogen in toluene and synthesize benzoic acid however i cannot find exact mechanism of reaction in

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web description treatment of an alkylbenzene with potassium permanganate results in oxidation to give the benzoic acid notes the position directly adjacent to an aromatic group is called the benzylic position the reaction only works if there is a hydrogen attached to the carbon examples

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web may 1 2004 the oxidation of toluene by aq KMnO_4 under hydrodynamic cavitation was taken as a model reaction and various parameters have been optimized the oxidation of toluene by aq KMnO_4 gives benzoic acid scheme 1 KMnO_4 in turn is reduced to MnO_2 the reaction did not give any other byproduct