PRINCIPLES RELATED TO PRACTICAL CHEMISTRY

SYNOPSIS

Detection of elements (Organic):

- Carbon and Hydrogen: The given organic compound is mixed with double the amount of pure and dry copper oxide the mixture is heated in hard glass tube.
- The CO₂ and H₂O produced due to combustion are tested by lime water and anhydrous copper sulphate.
- Lime water turns milky indicates the presence of carbon and anhydrous CuSO to blue indicates the presence of hydrogen.

$$C + 2CuO \longrightarrow Host \longrightarrow CO_2 + 2Cu$$

 $2H + CuO \longrightarrow H_2O + Cu$
 $CuSO_4 + 5H_2O \longrightarrow CuSO_4.5H_2O$
volute blue

$$Ca(OH)_2 + CO_2 \rightarrow CaCO_1 + H_2O$$
milky white

- Soda lime test (for nitrogen): When the organic compound is heated with sodalime in a test tube, the evolution of NH, indicates nitrogen CH, CONH, + NaOH CH, COONa + NH.
- Lassaigne's test: This is used to detect nitrogen, halogen and sulphur.
- Organic compound is fused with dry sodium in a fusion tube and fused mass after extraction with water is boiled and filtered. Filterate is called sodium extract (S.E) is used to detect elements (other than C and H).
- Sodium extract is made alkaline with a few drops of dil NaOH.
- Fusion with Na forms soluble salt (like NaCl, NaCN, etc) which can be easily detected.

Element	Sodium extract (S.E)	Confirmed test	Reactions
Nitropen	$Ma + C + N \xrightarrow{\Delta} MaCN$ $(\mathcal{E}.\mathcal{E})$	S.E + FeSO, + NaOH, boil and cool + FeCl, + conc.HCl Blue or green colour	FeSO ₄ + 2 NaOH \rightarrow Fe(OH) ₁ + Na ₂ SO ₂ Fe(OH) ₂ + 6 Na CN \rightarrow Na ₃ [Fe(CN) ₃] + 2 NaOH 3 Na ₄ [Fe(CN) ₃] + 4 FeCl ₂ \rightarrow Fe ₄ [Fe(CN) ₃] ₄ + 12 NaCI Pression blue
Sulphur	$2Na + S \xrightarrow{\Delta} Na_2S$ $(S.E)$	(I) S.E. + sodium nitro pruoside (A). A deep violet colour. (ii) S.E. + CH,COOH + (CH,COO), Pb Black ppt.	00 Na_S+ Na_BFetCNs, NOt ++ Na_BFetCNs, NOS: tAs deep violes 003 Na_S+ (CH_COO), Pt. CH_COOH Pas 4 = 2 CH_COONs black pps
Halogen	$Na + Cl \xrightarrow{\Delta} NaCl \atop (S.E)$	S.E. * HNO, * AgNO, (i) White ppt. soluble in aq NH, confirms CI (ii) Yellow ppt. partially soluble in aq NH, confirms Br. (iii) Yellow ppt insoluble in aq NH, confirms I.	NaX+AgNO, HNO, AgX 1 ppt AgCl +2NH,(aq) - [Ag(NH,),]Cl white ppt soluble
Nationers and sulphor together	Na+ C+ N+ S	As in test for nitrogen, instead of green or blue colour, blood red colour, blood red colourism confirms presence of N and S both.	3 NaCNS+ FeCl, = [Fe(CNS),] + 3 NaCl Blood red colour

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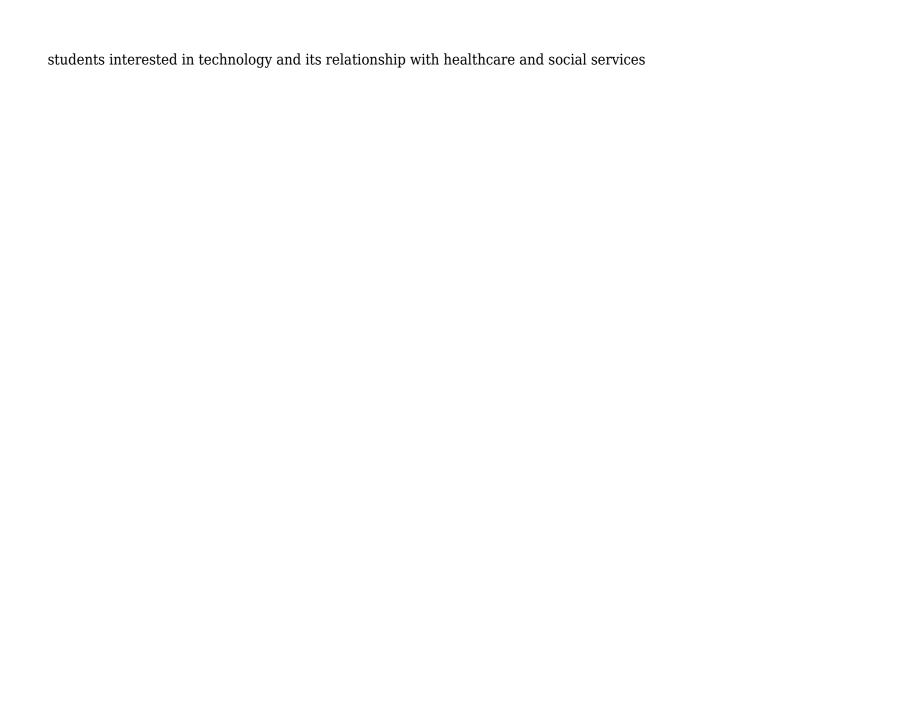
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