

PRINCIPLES RELATED TO PRACTICAL CHEMISTRY

SYNOPSIS

Detection of elements (Organic):

- **Carbon and Hydrogen:** The given organic compound is mixed with double the amount of pure and dry copper oxide the mixture is heated in hard glass tube.
- The CO_2 and H_2O produced due to combustion are tested by lime water and anhydrous copper sulphate.
- Lime water turns milky indicates the presence of carbon and anhydrous CuSO_4 to blue indicates the presence of hydrogen.



- **Soda lime test (for nitrogen):** When the organic compound is heated with sodalime in a test tube, the evolution of NH_3 indicates nitrogen
- $\text{CH}_3\text{CONH}_2 + \text{NaOH} \xrightarrow{\text{Heat}} \text{CH}_3\text{COONa} + \text{NH}_3$
- **Lassaigne's test:** This is used to detect nitrogen, halogen and sulphur.
- Organic compound is fused with dry sodium in a fusion tube and fused mass after extraction with water is boiled and filtered. Filtrate is called sodium extract (S.E) is used to detect elements (other than C and H).
- Sodium extract is made alkaline with a few drops of dil NaOH.
- Fusion with Na forms soluble salt (like NaCl, NaCN, etc) which can be easily detected.

Element	Sodium extract (S.E)	Confirmed test	Reactions
Nitrogen	$\text{Na} + \text{C} + \text{N} \xrightarrow{\Delta} \text{NaCN}$ (S.E)	S.E + FeSO_4 + NaOH , boil and cool + FeCl_3 + conc. HCl Blue or green colour	$\text{FeSO}_4 + 2\text{NaOH} \rightarrow \text{Fe(OH)}_2 + \text{Na}_2\text{SO}_4$ $\text{Fe(OH)}_2 + 6\text{NaCN} \rightarrow \text{Na}_4[\text{Fe(CN)}_6] + 2\text{NaOH}$ $3\text{Na}_4[\text{Fe(CN)}_6] + 4\text{FeCl}_3 \rightarrow \text{Fe}_3[\text{Fe(CN)}_6]_4 + 12\text{NaCl}$ Prussian blue
Sulphur	$2\text{Na} + \text{S} \xrightarrow{\Delta} \text{Na}_2\text{S}$ (S.E)	(i) S.E. + sodium nitro peroxide (A). A deep violet colour. (ii) S.E. + CH_3COOH + $(\text{CH}_3\text{COO})_2\text{Pb}$ Black ppt.	(i) $\text{Na}_2\text{S} + \text{Na}_2\text{Fe(CN)}_6 \cdot \text{NO} \rightarrow \text{Na}_4[\text{Fe(CN)}_6 \cdot \text{NO}]$ LAS deep violet (ii) $\text{Na}_2\text{S} + (\text{CH}_3\text{COO})_2\text{Pb} \xrightarrow{\text{CH}_3\text{COOH}} \text{PbS} \downarrow + 2\text{CH}_3\text{COONa}$ black ppt
Halogen	$\text{Na} + \text{Cl} \xrightarrow{\Delta} \text{NaCl}$ (S.E)	S.E. + HNO_3 + AgNO_3 (i) White ppt. soluble in aq. NH_3 , confirms Cl (ii) Yellow ppt. partially soluble in aq. NH_3 , confirms Br. (iii) Yellow ppt. insoluble in aq. NH_3 , confirms I.	$\text{NaX} + \text{AgNO}_3 \xrightarrow{\text{HNO}_3} \text{AgX} \downarrow$ ppt $\text{AgCl} + 2\text{NH}_3(\text{aq}) \rightarrow [\text{Ag(NH}_3)_2]\text{Cl}$ white ppt. soluble
Nitrogen and sulphur together	$\text{Na} + \text{C} + \text{N} + \text{S} \xrightarrow{\Delta} \text{NaCNS}$ (S.E) (Sodium thiocyanate)	As in test for nitrogen, instead of green or blue colour, blood red coloration confirms presence of N and S both.	$3\text{NaCNS} + \text{FeCl}_3 \rightarrow [\text{Fe(CNS)}_3] + 3\text{NaCl}$ Blood red colour

Principles Of Chemistry A Series Ofs In Chemistry

Louis Nicolas Vauquelin



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