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**Albert A Gayle**



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$$\text{Pb(NO}_3)_2(\text{aq}) + 2\text{HCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{HNO}_3$$

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web  $\text{Cu}(\text{ClO}_4)_2$  is soluble all perchlorates are soluble while  $\text{PbSO}_4$  is insoluble sulfates are generally soluble except for  $\text{Pb}^{2+}$   $\text{Ba}^{2+}$   $\text{Hg}_2^{2+}$  and  $\text{Sr}^{2+}$  the net ionic equation is

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