

Lose Weight Naturally



Losing Weight Naturally Wi Help Yourself Guides

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Harris**



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web limiting reactant theoretical and percentage yields key $\text{K}_2\text{S} + 2\text{H}_2\text{O} + 4\text{KOH} \rightarrow 3\text{O}_2 + 4\text{K}_2\text{O}$ if a reaction vessel contains 0.1 mol K_2O and 0.1 mol H_2O what is the limiting reactant?

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web limiting reactant: reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated. Percent yield: measure of the efficiency of a reaction expressed as a percentage of the theoretical yield.

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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation: $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield. So percent yield = $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$. Problems 6 & 8 are similar to problem 5.

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produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have 0.216232 grams/mole 21.3 grams / 0.66 moles = 32 g

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web a write a balanced equation for the reaction $Al + O_2 \rightarrow Al_2O_3$ 4 Al + 3 $O_2 \rightarrow 2 Al_2O_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x 1 mol Al / x 2 mol Al_2O_3 x 101.964 g Al_2O_3 / 302.3 g Al_2O_3 = 26.982 g Al 4 mol Al / 1 mol Al_2O_3 c determine the percent yield yield actual x 100 26.00 g / x 100 = 86.01% theoretical 302.3 g

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web 7 Na + 2 C + 2 O₂ → 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H₂ are left over 13

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield / actual yield 100 based on this definition we would expect a percent yield

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