

# CHAPTER 1

## MATTER AND MEASUREMENT

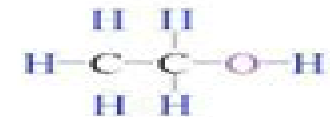
Chemistry: deals with composition, structure, and reactions of matter

- **Matter** – has mass and occupies space
  - Mass – measures quantity of matter
  - Weight – result of gravitational attraction between matter
    - Weight and mass are NOT the same
- **Composition** – what matter is made of and how much of each component is present
  - Expressed by...
    - Weight/mass
    - Volume
    - Percent
    - # of moles
    - # of atoms
  - *Macroscopic* – can be seen and weighed
  - *Submicroscopic level* – number and types of atoms
    - *Atoms* = simple units of matter
    - *Molecules* = combinations of atoms
      - Don't separate when dissolved like compounds → ionic compounds don't exist as molecules and separate when dissolved
    - *Qualitative* = ethanol consists of carbon, hydrogen and oxygen
    - *Quantitative* = ethanol =  $C_2H_6O$
- **Structure** – arrangement of components and how they are bonded
  - Different arrangements of same combination of components = isomers
  - Structures rotate freely
  - Dashed lines = behind                      dark triangles = forward
- **Reactions** – changes in composition and structure
  - What products?
  - How much of product? (stoich)
  - How fast change occurs?
  - What energy changes accompany rxn?



- exothermic reaction → produces energy over a long period of time

Ethanol  
 $C_2H_6O$   
or  
 $C_2H_5OH$



## Scientific Method

- Experiment (record observations) → numerical + observational
  - Careful recordings and analysis of data
    - Controlled conditions so we pinpoint what is causing another thing to happen
  - Reproducible
- Draw a Conclusion – **LAW**
  - Concise statement about a basic relationship or regularity of nature drawn from observations

# Notes On General Chemistry

**Zhe Xu**



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