

MÉMOIRES

....

Dialogue métallurgique (suite) ⁽¹⁾

par M. Albert SAUVEUR

(Traduit par M. Henri AUBERT)

E. — Je vous serais reconnaissant de me l'expliquer. Je suis curieux de savoir en quoi elle consiste.

P. — Cette curiosité vous fait honneur. On croit que le cuivre présent dans l'alliage se combine avec l'aluminium pour former le composé CuAl^2 .

E. — Est-ce un véritable composé chimique?

P. — Oui et non. Si la valence doit être rigoureusement respectée quand les éléments forment des corps composés, je crains que CuAl^2 ne puisse être considéré comme une combinaison. Si l'aluminium possède la valence 3, comme l'existence du composé AlCl^3 semble l'indiquer, le cuivre devrait avoir la valence 6 quand il se combine avec l'aluminium pour former CuAl^2 . Mais la valence du cuivre est soit 1, comme dans le composé CuCl , soit 2, comme dans le composé CuCl^2 (Jeffries, « Science of Metals»). Bien que ces composés intermétalliques ne diffèrent pas des métaux composants autant que les combinaisons chimiques ordinaires des éléments qui les constituent, la différence est cependant, dans de nombreux cas, très notable. C'est ce qui arrive d'ordinaire en ce qui concerne la dureté du composé qui est en général beaucoup plus grande. La dureté Brinell du cuivre, par exemple, est 40 et celle de l'aluminium 25, tandis que la dureté Brinell du composé CuAl^2 est 300. Une telle dureté est fréquemment accompagnée de fragilité. En d'autres termes, deux métaux doux et ductiles donnent souvent naissance en se combinant à un composé dur et fragile.

E. — Qu'est-ce que la dureté Brinell?

P. — J'attendais cette question et j'y répondrai bientôt.

E. — Mais avez-vous toujours l'intention de répondre à ma question initiale?

P. — Certainement, lorsque le moment sera venu, j'essaierai d'y répondre.

E. — Puisse ce moment arriver!

P. — Ayez confiance et en attendant retournons au vieillissement du duralumin. On suppose que le composé CuAl^2 est dissous dans l'aluminium et forme avec lui une solution solide.

E. — Vous m'avez dit — je m'en souviens — qu'il existait deux sortes de solutions solides, les solutions de substitution et les solutions d'insertion. À quel type les solutions solides de cuivre et d'aluminium appartiennent-elles?

P. — Ce sont des solutions de substitution. Les spécialistes nous disent cependant qu'il n'existe généralement pas de place dans les réseaux cristallins pour les molécules étrangères. Celles-ci doivent donc se dissocier avant d'y pénétrer. Admettant que le cuivre est en solution solide dans l'aluminium, nous avons des preuves concluantes que sa solubilité diminue quand la température baisse. Elle est bien moindre à la température ordinaire qu'à 500 degrés. À cette dernière température, l'aluminium peut retenir 5 % de cuivre en solution solide, tandis qu'à la température ordinaire, il ne peut en retenir que 2 %. Au cours d'un refroidissement lent de 500 degrés C à la température ordinaire, 3 % de cuivre sous la forme du composé CuAl^2 doivent donc être rejetés hors de la solution solide pour

(1) Voir Revue de Métallurgie, t. XXXIV, n° 3, mars 1937, p. 205, et n° 4, avril 1937, p. 253.

Metallurgical Reminiscences And Dialogue

Ian Pickup



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Table of Contents Metallurgical Reminiscences And Dialogue

1. Understanding the eBook Metallurgical Reminiscences And Dialogue
 - The Rise of Digital Reading Metallurgical Reminiscences And Dialogue
 - Advantages of eBooks Over Traditional Books
2. Identifying Metallurgical Reminiscences And Dialogue
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Metallurgical Reminiscences And Dialogue
 - User-Friendly Interface
4. Exploring eBook Recommendations from Metallurgical Reminiscences And Dialogue
 - Personalized Recommendations
 - Metallurgical Reminiscences And Dialogue User Reviews and Ratings
 - Metallurgical Reminiscences And Dialogue and Bestseller Lists
5. Accessing Metallurgical Reminiscences And Dialogue Free and Paid eBooks
 - Metallurgical Reminiscences And Dialogue Public Domain eBooks
 - Metallurgical Reminiscences And Dialogue eBook Subscription Services
 - Metallurgical Reminiscences And Dialogue Budget-Friendly Options

6. Navigating Metallurgical Reminiscences And Dialogue eBook Formats
 - ePub, PDF, MOBI, and More
 - Metallurgical Reminiscences And Dialogue Compatibility with Devices
 - Metallurgical Reminiscences And Dialogue Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Metallurgical Reminiscences And Dialogue
 - Highlighting and Note-Taking Metallurgical Reminiscences And Dialogue
 - Interactive Elements Metallurgical Reminiscences And Dialogue
8. Staying Engaged with Metallurgical Reminiscences And Dialogue
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Metallurgical Reminiscences And Dialogue
9. Balancing eBooks and Physical Books Metallurgical Reminiscences And Dialogue
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Metallurgical Reminiscences And Dialogue
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Metallurgical Reminiscences And Dialogue
 - Setting Reading Goals Metallurgical Reminiscences And Dialogue
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Metallurgical Reminiscences And Dialogue
 - Fact-Checking eBook Content of Metallurgical Reminiscences And Dialogue
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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