Chapitre 2 Analyser un système chimique par des méthodes physiques

1) Analyse par mesure d'absorbance

L'absorbance est une grandeur physique caractérisant la capacité d'une solution à absorber la lumière.

Loi de Beer-Lambert :

A une longueur d'onde donnée, l'absorbance A_{λ} d'une espèce chimique en solution diluée est proportionnelle à la concentration molaire C de cette espèce et à l'épaisseur C de la solution traversée :

absorbance de la solution (sans unité)

 $A_{\lambda} = E_{\lambda} \times l \times C = k \times C$

concentration molaire de l'espèce chimique (mol.L⁻¹)

coefficient d'extinction molaire (Lmol-1.cm-1)

Longueur de la cuve (cm)

La loi de Beer-Lambert n'est valable que pour des solutions diluées dont l'absorbance est inférieure à 2.

2) Analyse par mesure de conductance activité expérimentale page 43

a) Les solutions conductrices

Pour qu'une solution conduise le courant électrique, il faut qu'elle contienne des lons

Dans une solution ionique, les cations se déplacent dans le sens conventionnel du courant alors que les anions se déplacent dans le sens



La conductance d'une portion de solution comprise entre 2 électrodes permet d'estimer le caractère conducteur d'une solution. C'est l'inverse de la résistance.

G : conductance en Siemens (S) $G = \frac{1}{R} = \frac{I}{U}$ I : intensité du courant (A)

R : résistance en Siemens (Ω) U : tension entre les électrodes (V)

c) Conductivité d'une solution

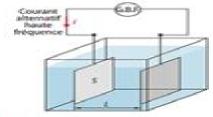
On constate expérimentalement, que la conductance est proportionnelle à :

la surface des électrodes

l'inverse de la longueur entre les électrodes.

Le coefficient de proportionnalité s'appelle la conductivité de la solution ;

$$G = \sigma \times \frac{E}{L}$$
 surface des électrodes (m²) conductivité de la solution (S.m⁻¹) distance entre les électrodes (m)



On constate expérimentalement que la conductance G augmente avec la température de la solution ainsi qu'avec la concentration en soluté apporté de la solution et dépend de l'état de la surface des électrodes (propre, lisse, ...).

La conductivité d'une solution aqueuse est une grandeur qui représente la capacité de cette solution à conduire le courant électrique. Elle se note σ et s'exprime en siemens par mètre

conductivité (S)
$$\sigma = G \times \frac{L}{S} = k \times G + \text{conductance (S.m-1)}$$

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