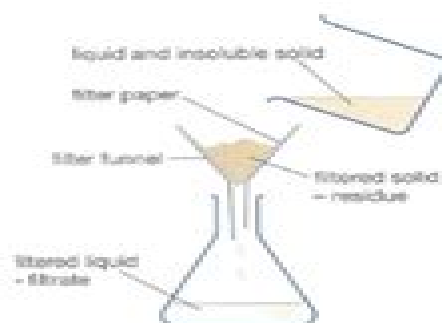


Separation techniques study guide

1. Filtration- (used to separate a solid in liquid mixture where the solid does not dissolve in the solvent and forms a heterogeneous solution. Example: a mixture of sand and water or soil and water)



The part of the mixture that stays behind in the filter paper is called the residue.

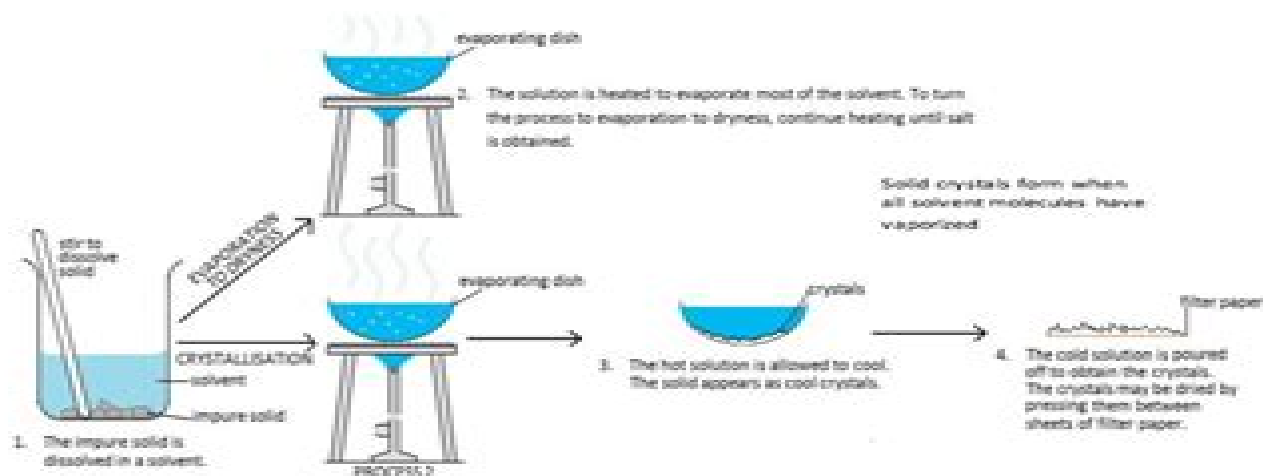
The part of the mixture that passes through the filter paper is called the filtrate.



Filtration: An undissolved/insoluble solid can be separated from a solution by passing it through a piece of filter paper in a filter funnel. The solution which passes through the filter paper is called the filtrate. The solid that stays on the filter paper is called the residue. This residue should then be washed with distilled water to remove any solution between its solid particles.

It operates on the principle of exploiting **particle size differences**. The filter paper acts as a **selectively permeable membrane** as it allows smaller sized substances to pass through while not allowing larger sized particles to do so.
Example: mixture of sand and water or soil and water

2. Crystallization



Crystallization is a process used to separate a solid in liquid mixture where the solid is soluble in the solvent and dissolves in it and forms a homogeneous solution. Example: A mixture of NaCl and water.

It exploits the principle of boiling point differences.

A crystalline solid can be extracted from a solution by heating the solution. The solvent is evaporated until the crystallization point of the solid is reached. You can know if the crystallization point is reached by placing a drop of the solution on a cold surface from time to time to see if crystals form. When it does leave the solution to cool on a cold surface until crystals form fully as we did in the lab.

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