THE NATURE OF THE ATOM

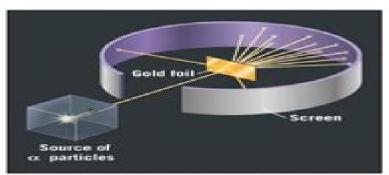
Rutherford Scattering and the Nuclear Atom

Early Greek models of atoms described them as tiny indivisible spheres. In fact atomos, the Greek word for indivisible, is the origin of the word atom.

The plum pudding model was an improvement proposed after electrons were discovered. The idea was that electrons were embedded in a sphere of positive matter like plums in a pudding.

This model was replaced with the planetary model proposed by Ernest Rutherford. He directed a beam of alpha particles at a very thin sheet of gold foil.

Most of the alpha particles passed through undeflected as expected. Some, however, were deflected up to 180 degrees. The only explanation was that the positive charge of each atom must be concentrated in a relatively small volume which was named the nucleus.



The amount of empty space in an atom is difficult to imagine. It may help to imagine an atom the size of our solar system. If the nucleus of an atom were about the size of the sun, the

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