

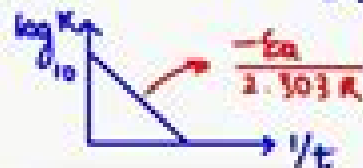
Temperature

$$A = \frac{K_2 + 10}{K_1}$$

$$\frac{k_2}{k_1} = \frac{K_2}{K_1}$$

Arrhenius equation

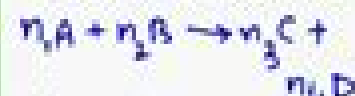
$$\log_{10} K = \log_{10} A - \frac{E_a}{2.303RT}$$



Collision theory

- reactants must have threshold Energy
- Proper orientation

Rate of reaction



$$-\frac{1}{n_1} \frac{d[A]}{dt} = -\frac{1}{n_2} \frac{d[B]}{dt}$$

$$= \frac{1}{n_3} \frac{d[C]}{dt} = \frac{1}{n_4} \frac{d[D]}{dt}$$

Rate law  $\rightarrow K[A]^x[B]^y$   
order  $\rightarrow x+y$

Nth order reaction

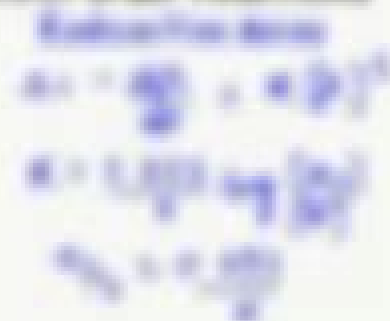
$$n \neq 1 \rightarrow k t^{n-1} \left[ \frac{1}{(n-1)} \left( \frac{1}{[A]^{n-1}} - \frac{1}{[A]_0^{n-1}} \right) \right]$$

$$t_{1/2} = \frac{1}{[A]_0^{n-1}}$$

# Chemical kinetics

Molecularity

First order reactions



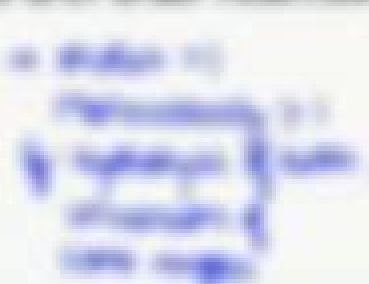
Zero order reaction



Elementary reactions

- forming order is equal
- molecularity  $\leq$  order

Overall first order reaction



# Research In Chemical Kinetics

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**Air Force Research Objectives** United States. Air Force. Office of Aerospace Research, **Summary Report on the Workshop on High Temperature Chemical Kinetics, Applications to Combustion Research** National Measurement Laboratory (U.S.), David Garvin, 1978

*Chemical Kinetics* Luis Arnaut, 2021-06-22

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