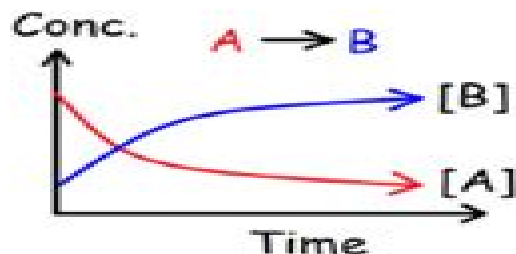


## Chemical Kinetics:



**Instantaneous Rate of Appearance:**

$$\text{Rate} = + \frac{d[B]}{dt}$$

**Method of Initial Rates:**

Trial	[A]	[B]	[C]	I. Rate
1	0.10 M	0.10 M	0.10 M	0.20 M/s
2	0.20 M	0.10 M	0.10 M	0.40 M/s
3	0.10 M	0.20 M	0.10 M	0.80 M/s
4	0.10 M	0.10 M	0.20 M	0.20 M/s

**Rate Constant k:**

$$k = \frac{\text{Rate}}{[A]^x[B]^y[C]^z}$$

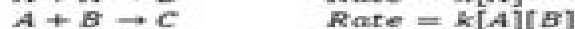
**Units of k:**  $M^{1-n} t^{-1}$  or  $(mol)^{1-n} (L)^{n-1} t^{-1}$

**Note:**  $M = mol \cdot L^{-1}$  and  $t \rightarrow s, min, hr, days$

**Unimolecular:**



**Bimolecular:**



**Termolecular:**



**Average Rate of Appearance:**  $A \rightarrow B$

$$\text{Rate} = + \frac{\Delta[B]}{\Delta t} = \frac{[B]_f - [B]_o}{t_f - t_o}$$

**Average Rate of Disappearance:**

$$\text{Rate} = - \frac{\Delta[A]}{\Delta t}$$

**Rate of a Chemical Reaction:**



$$\text{Rate} = - \frac{1}{2} \frac{\Delta[A]}{\Delta t} = - \frac{1}{3} \frac{\Delta[B]}{\Delta t} = + \frac{1}{4} \frac{\Delta[C]}{\Delta t} = + \frac{1}{5} \frac{\Delta[D]}{\Delta t}$$

**Differential Rate Law Expression:**

$$\text{Rate} = k[A]^x[B]^y[C]^z$$

**Finding The Order of a Reactant:**

$$x = \frac{\log \left( \frac{\text{Rate 2}}{\text{Rate 1}} \right)}{\log \left( \frac{[A]_2}{[A]_1} \right)} \quad y = \frac{\log \left( \frac{\text{Rate 3}}{\text{Rate 1}} \right)}{\log \left( \frac{[B]_3}{[B]_1} \right)}$$

**Overall order of the reaction:**

$$\text{Order} = x + y + z$$

**Factors Affecting the Rate of a Reaction:**

1. Temperature
2. Concentration
3. Catalyst
4. Surface Area
5. The Nature of the Reactants

**Reaction Mechanisms:**



**Note:** Catalyst  $\rightarrow B$  Intermediate  $\rightarrow D$

# Research In Chemical Kinetics V 1 Research In Chemical Kinetics

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Development Administration**



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