

NINTH EDITION

THE SCOTT, FRESHMAN HANDBOOK FOR WRITERS

JOHN RUSZKIEWICZ
University of Texas at Austin

CHRISTY FRIEND
University of South Carolina

DANIEL E. SEWARD

MAXINE HAIRSTON
Late, University of Texas at Austin



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web apr 7 2014 molarity moles of solute liters of solution 8 4 2 2 a first convert 250 ml to liters 250 1000 0 25 then calculate molarity 5 moles 0 25 liters 20 m 3 c a solution with molarity 2 requires 2 m of n a oh per liter so 4 x 2 8 m 4 a a solution of molarity 1 5 m requires 1 5 mol of na to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture mmix can be calculated using the following formula

$$m_{mix} = \frac{m_1 + m_2 + m_3}{V_1 + V_2 + V_3}$$
 1 5 x 480 1 2 x 520 480 520 720 624 1000 1 344 m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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web determine the molarity for each of the following solutions 0 444 mol of cocl 2 in 0 654 l of solution 98 0 g of phosphoric acid h 3 po 4 in 1 00 l of solution 0 2074 g of calcium hydroxide ca oh 2 in 40 00 ml of solution 10 5 kg

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