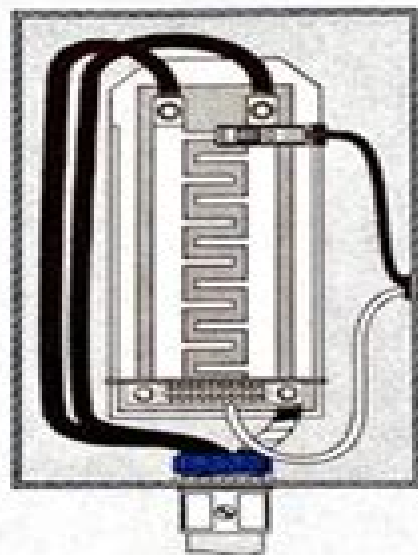
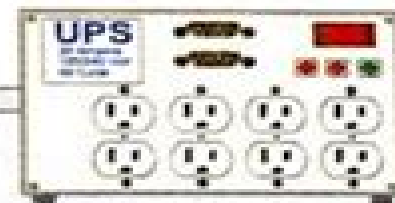




Voltage Drop - Single Phase (Ohm's Law Method) $VD = I \times R$

$D = 100$ Feet

No. 12 THHN



Load
16 Amps

Calculating voltage drop using Ohm's Law.

Ohm's Law
 $E = I \times R$

To use this method, "I" and "R" must be known. "I" is given as 16 amperes. "R" or resistance of the conductors must be calculated.

$$\text{Resistance} = \left(\frac{\text{Table 9 ohms}}{1,000 \text{ feet}} \right) \times \text{number of feet}$$

Chapter 9 Table 9:

Resistance of No. 12 = 2 ohms per 1,000 feet

Number of Feet: 100 feet x 2 conductors = 200 feet

$$\text{Resistance} = \left(\frac{2 \text{ ohms}}{1,000 \text{ feet}} \right) \times 200 \text{ feet} = 0.4 \text{ ohms}$$

$$VD = I \times R = 16 \text{ amperes} \times 0.4 \text{ ohms} = 6.4 \text{ volts}$$

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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2 \text{ kJ/mol rxn}$ write equations to similarly represent the dissolving process for

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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