

## Part I - Multiple Choice (45 pts, 3 pts each).

1) Which one of the following is not a state function?

- A) enthalpy H  
 B) heat  $q$   
 C) temperature T  
 D) pressure P  
 E) internal energy U

Use the table of specific heats to answer questions 2 and 3

2) Which of the following substances would need the largest amount of heat to change the temperature of 1 gram by  $10^\circ\text{C}$ ?

- A) Aluminum, Al  
 B) Copper, Cu  
 C) Gold, Au  
 D) Lead, Pb  
 E) Silver, Ag

| Substance | Specific Heat                     |
|-----------|-----------------------------------|
| Aluminum  | $0.900 \text{ J/g}^\circ\text{C}$ |
| Copper    | $0.386 \text{ J/g}^\circ\text{C}$ |
| Gold      | $0.126 \text{ J/g}^\circ\text{C}$ |
| Lead      | $0.128 \text{ J/g}^\circ\text{C}$ |
| Silver    | $0.233 \text{ J/g}^\circ\text{C}$ |

$$q = cm\Delta T$$

$\downarrow$   $\downarrow$   $\downarrow$   
 $10 \text{ g}$   $10^\circ\text{C}$

3) Which would need the largest amount of heat to change the temperature of 1 mole by  $10^\circ\text{C}$ ?

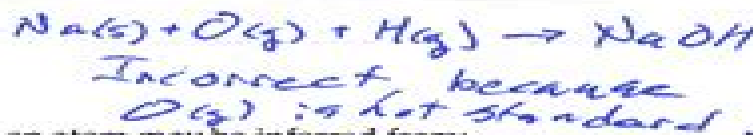
- A) Aluminum, Al  
 B) Copper, Cu  
 C) Gold, Au  
 D) Lead, Pb  
 E) Silver, Ag

$$q = C_m \Delta T = M C_m \Delta T$$

$\downarrow$   $\downarrow$   $\downarrow$   
 $26.98 \text{ g/mol}$   $0.900 \text{ J/g}^\circ\text{C}$   $10 \text{ mole}$

4) The enthalpy of formation,  $\Delta H_f^\circ$ , of NaOH (s) is the same as the  $\Delta H_r^\circ$  for which of these reactions?

- A)  $\text{Na (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaOH (s)} + \frac{1}{2} \text{H}_2 \text{ (g)}$   
 B)  $\text{Na (s)} + \frac{1}{2} \text{O}_2 \text{ (g)} + \frac{1}{2} \text{H}_2 \text{ (g)} \rightarrow \text{NaOH (s)}$   
 C)  $\text{Na (s)} + \frac{1}{2} \text{H}_2\text{O}_2 \text{ (l)} \rightarrow \text{NaOH (s)}$   
 D)  $\text{Na}^+ \text{ (aq)} + \text{OH}^- \text{ (aq)} \rightarrow \text{NaOH (s)}$   
 E)  $\text{N (g)} + \text{a} + \text{O (g)} + \text{H}^- \text{ (aq)} \rightarrow \text{NaOH (s)}$



5) The existence of discrete (quantized) energy levels in an atom may be inferred from:

- A) diffraction of electrons by crystals  
 B) atomic line spectra  
 C) visible spectrum  
 D) experiments on the photoelectric effect  
 E) X-ray diffraction by crystals

6) Arrange the following in order of increasing energy:

Ultra-violet (UV) radiation, visible radiation, gamma-rays, and infra-red (IR) radiation.

IR                      visible                      UV                      gamma  
 Lowest                                                                                     Highest

# Sg Sm T A Gen Chem 3 E The Ori

**Michael Hoelscher**



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