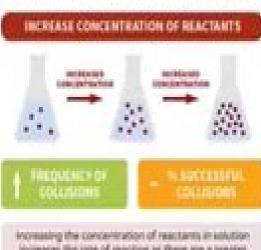
REACTION KINETICS

MAKING CHEMICAL REACTIONS HAPPEN FASTER

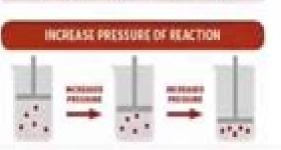
There are a number of different things that we can change to make a chemical reaction faster. Here, we explain the concept of collision theory, and how it can be used to explain the effects of five different factors on the rate of a chemical reaction.

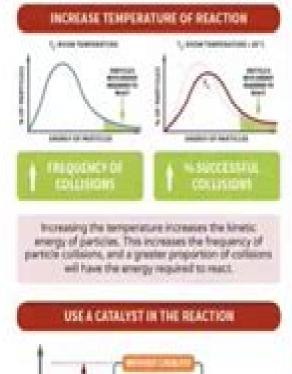
Colleges theory states that, for a reaction to occur, particles must suited with the correct orientation and such sufficient energy for a reaction to occur. Different factors affect the rate of the reaction by affecting the frequency of particle collisions, and/or the proportion of collisions that have enough energy to react.

MICHEASE SUPPLIE AREA OF REACTANTS



increasing the concentration of reactants in solution increases the rate of reaction as there are a greater number of particles available to react. This increases the frequency of collopons between particles.





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Reaction Kinetics In Heterogeneous Chemical Systems

Mike Jess

Reaction Kinetics In Heterogeneous Chemical Systems:

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