

CHEM 1212
Pre-Lab Instructions/Lab Report Template
Effect of Temperature on the Solubility of a Salt

Complete all of the following in directly on this form or in your lab notebook.

Provide the title **Effect of Temperature on Solubility of a Salt Lab**

Provide an objective for the lab.

The objective of this lab is to understand how temperature will affect how much KNO_3 will dissolve in different amounts of water and effectively plot the results on a data sheet.

Safety

- Provide a molecular formula, a molar mass, a melting point, and safety-hazard information for potassium nitrate.
 - KNO_3 , 101.10g/mol, 334 °C, Oxidizer and health hazard
- Provide a molecular formula, a molar mass, a density, a boiling point, and safety-hazard information for water.
 - H_2O , 18.02 g/mol, 997 kg.m⁻³, 100 °C, no safety concerns

Summarize the procedure:

Obtain the different amounts of KNO_3 and add 5.0 mL of water to each. Set up the temperature probe on LabQuest. Fill the large beaker with water and place on the hot plate. Use the temperature probe to measure the temperature of the solution on the ring stand. Once the KNO_3 is dissolved take out of the warm bath and place on another ring stand while waiting for crystallization to occur. While this is occurring, continuously stir the solution with the temperature probe. If no crystallization occurs, place in ice bath after the temperature reaches room temperature. Once the temperatures are recorded return the tubes to the rack and clean probe. Repeat these steps for the other test tubes. After completing the other test tubes stop data collection and discard contents in the fume hood then use excel to create a graph.

List any equations that you plan to use in this lab.

Grams of KNO_3 / 5 grams of H_2O = X grams of KNO_3 / 100 grams of H_2O

Complete in Lab:

DATA:

Test tube number	Amount of KNO_3 used per 5 mL H_2O	Amount of KNO_3 used per 100g H_2O	Temperature at Crystallization
1	2.0	80	21.2
2	4.0	80	44.0
3	6.0	120	65.3
4	8.0	160	71.5

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