

Fig. 2.1 shows the general configuration of a thermodynamic system at equilibrium at temperature  $T$ . The internal energy of the gas is the important macroscopic (state) variable.

Assume  $P_1$  is small, negligible, and  $P_2$  constant, and hence

$$\frac{dW}{dt} = \frac{P_2}{A} \frac{dV}{dt}$$

Assume  $P_1$  is infinitesimally small,  $P_2$  is 100 kPa,  $A$  is 0.01 m<sup>2</sup>, and  $\frac{dV}{dt}$  is 0.01 m<sup>3</sup>/s.  $P_2$  is 100 kPa,  $A$  is 0.01 m<sup>2</sup>, and  $\frac{dV}{dt}$  is 0.01 m<sup>3</sup>/s.

Then

$$P_{\text{ext}} = \frac{P_1 + P_2}{2} = \frac{\text{Pressure of gas} + P_2}{2} = P_2 \text{ is small}$$

Fig. 2.1 shows the general configuration of a thermodynamic system at equilibrium at temperature  $T$ . The internal energy of the gas is the important macroscopic (state) variable. The internal energy of the gas is the important macroscopic (state) variable. The internal energy of the gas is the important macroscopic (state) variable.

$$\frac{dW}{dt} = \frac{P_2}{A} \frac{dV}{dt} = \frac{P_2}{A} \frac{dV}{dt} = \frac{P_2}{A} \frac{dV}{dt}$$

$$\text{Internal energy of gas} = \frac{P_2}{A} \frac{dV}{dt} = P_2 \frac{dV}{dt}$$

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