

3. В сосуд, содержащий воду массой $m_1 = 200$ г при температуре $t_1^o = 20^o\text{C}$, опустили медное тело массой $m_2 = 100$ г, имеющее температуру $t_2^o = 120^o\text{C}$. Какая общая температура установится в сосуде? Нагреванием сосуда пренебречь. Удельные теплоёмкости воды и меди равны соответственно $c_1 = 4200$ Дж/(кг· 1^oC), $c_2 = 380$ Дж/(кг· 1^oC).

Решение. Определим направление теплообмена в системе "холодная вода + горячее медное тело": медное тело отдаёт тепло; холодная вода получает тепло.

$Q_1 = m_1 c_1 (t^o - t_1^o)$ – количество теплоты, полученное холодной водой.

$Q_2 = m_2 c_2 (t_2^o - t^o)$ – количество теплоты, отданное горячим медным телом.

В соответствии с уравнением теплового баланса (всеми потерями энергии пренебрегаем): $Q_1 = Q_2$.

Раскроем полученное равенство:

$$m_1 c_1 (t^o - t_1^o) = m_2 c_2 (t_2^o - t^o).$$

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