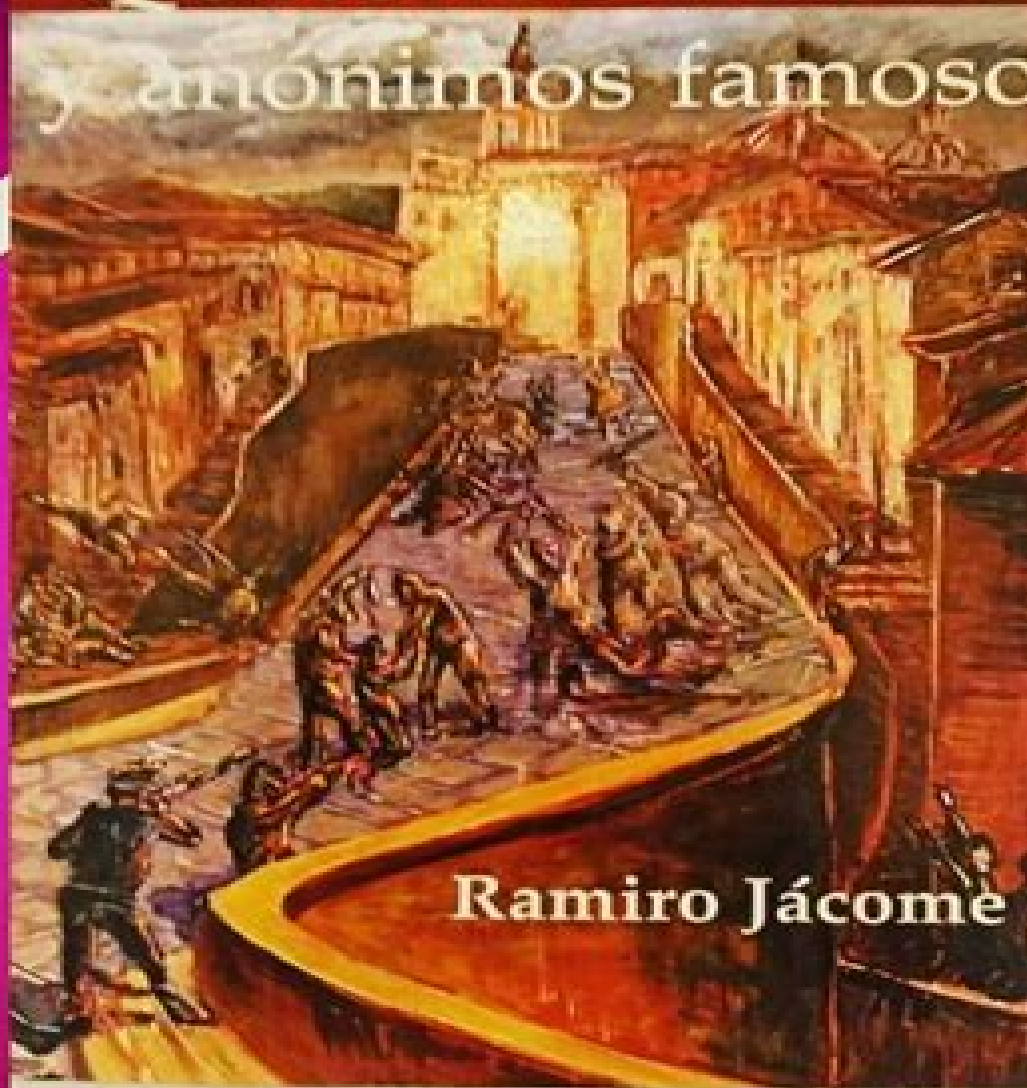


Quiteños ilustres y anónimos famosos



Ramiro Jácome

Quitenos Ilustres Y Anonimos Famosos Ramiro Jacome

Jorge Enrique Adoum



Quitenos Ilustres Y Anonimos Famosos Ramiro Jacome:

Quiteños ilustres y anónimos famosos Ramiro Jácome Durango, 2001 **Latin American Artists' Signatures and Monograms** John Castagno, 1997 Each entry includes if applicable nationality birth and death dates key letters to a listing of reference sources and the most recent auction citings if any of the last twenty years Nearly 4 800 artists and 9 000 signatures **Bibliographic Guide to Art and Architecture** New York Public Library. Art and Architecture Division, 1980

Quién es quién en Quito, 1997 **Vistazo**, 2002 En la tierra, Quito-- Jorge Enrique Adoum, 2004 Critic and historian Len n O a states in the introduction that the art genre Quito exists as a series of visual and artistic codes that have survived over time This edition is an anthology of the vast production of paintings drawings and prints inspired in the ancient city by renown local and foreign artists active between the 18th century to present days Rafael Salas Victor Mideros Joaqu n Pinto Carlos Kohn Sergio Guarderas Osvaldo Guayasam n Luis Moscoso Galo Galecio Pancho Fierro C sar Tacco Eduardo Kingman Luigi Casadio Paul Bar Ramiro J come Araceli Gilbert Gonzalo Endara Osvaldo Viteri Nicol s Svistoonoff Gilberto Almeida C ndido Portinari Frederick Edwin Church and Ruby Larrea amongst others Exhibition curated by Len n O a **Espejo, el ilustrado** Andrés Landázuri, 2011 *Quito integral* Fernando Rohn Bautista, 2008 **La Galería** Galería (Quito, Ecuador), Lenín Oña, 1997 **El Palacio de Carondelet** Jorge Salvador Lara, 1996 *Códigos arcanos* José Unda, 2006

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para

directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no₂⁺ which acts as the electrophilic species

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