

Chapter # 03

➔ Atomic Structure

ATOM:

Atom has three fundamental particles which are these:

I. ELECTRON:

Electron carries negative charge the charge on Electron is $1.6 \times 10^{-19} \text{C}$ and the mass of Electron is $9.1094 \times 10^{-31} \text{ Kg}$ or 0.0005 a. m. u.

II. PROTON:

Proton carries positive charge the charge on Proton is $1.6 \times 10^{-19} \text{C}$ and the mass of Proton is $1.6726 \times 10^{-27} \text{ Kg}$ or 1.0073 a. m. u.

III. NEUTRON:

Neutron is a neutral particle.

- Crook's tube experiment shows the presence of electron and protons in the atoms.
- In the Crook's tube experiment rays coming out from cathode are called cathode rays. These are negatively charged rays and consist of Electron.
- Neutron was discovered by James Chadwick.
- According to Rutherford, mass of atom is concentrated at its center called nucleus.
- The central nucleus of an atom is very small. It contains positively charged protons and neutrons.
- which have no charge.
- Goldstein discovered positive canal rays (protons). Canal rays travel in discharge tube from anode towards cathode.

RADIOACTIVITY:

The spontaneous decay of unstable atom with the emission of invisible radiation is called as Radioactivity, discovered by Henri Becquerel.

RADIOACTIVE ELEMENTS GIVEN OFF α , β AND γ RAYS.

- Alpha (α) particles are positively charged.
- Beta (β) particles are negatively charged.
- γ — Rays are neutral.

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