

MÉMOIRES

....

Dialogue métallurgique (suite) ⁽¹⁾

par M. Albert SAUVEUR

(Traduit par M. Henri AUBERT)

E. — Je vous serais reconnaissant de me l'expliquer. Je suis curieux de savoir en quoi elle consiste.

P. — Cette curiosité vous fait honneur. On croit que le cuivre présent dans l'alliage se combine avec l'aluminium pour former le composé CuAl^2 .

E. — Est-ce un véritable composé chimique?

P. — Oui et non. Si la valence doit être rigoureusement respectée quand les éléments forment des corps composés, je crains que CuAl^2 ne puisse être considéré comme une combinaison. Si l'aluminium possède la valence 3, comme l'existence du composé AlCl^3 semble l'indiquer, le cuivre devrait avoir la valence 6 quand il se combine avec l'aluminium pour former CuAl^2 . Mais la valence du cuivre est soit 1, comme dans le composé CuCl , soit 2, comme dans le composé CuCl^2 (Jeffries, « Science of Metals»). Bien que ces composés intermétalliques ne diffèrent pas des métaux composants autant que les combinaisons chimiques ordinaires des éléments qui les constituent, la différence est cependant, dans de nombreux cas, très notable. C'est ce qui arrive d'ordinaire en ce qui concerne la dureté du composé qui est en général beaucoup plus grande. La dureté Brinell du cuivre, par exemple, est 40 et celle de l'aluminium 25, tandis que la dureté Brinell du composé CuAl^2 est 300. Une telle dureté est fréquemment accompagnée de fragilité. En d'autres termes, deux métaux doux et ductiles donnent souvent naissance en se combinant à un composé dur et fragile.

E. — Qu'est-ce que la dureté Brinell?

P. — J'attendais cette question et j'y répondrai bientôt.

E. — Mais avez-vous toujours l'intention de répondre à ma question initiale?

P. — Certainement, lorsque le moment sera venu, j'essaierai d'y répondre.

E. — Puisse ce moment arriver!

P. — Ayez confiance et en attendant retournons au vieillissement du duralumin. On suppose que le composé CuAl^2 est dissous dans l'aluminium et forme avec lui une solution solide.

E. — Vous m'avez dit — je m'en souviens — qu'il existait deux sortes de solutions solides, les solutions de substitution et les solutions d'insertion. À quel type les solutions solides de cuivre et d'aluminium appartiennent-elles?

P. — Ce sont des solutions de substitution. Les spécialistes nous disent cependant qu'il n'existe généralement pas de place dans les réseaux cristallins pour les molécules étrangères. Celles-ci doivent donc se dissocier avant d'y pénétrer. Admettant que le cuivre est en solution solide dans l'aluminium, nous avons des preuves concluantes que sa solubilité diminue quand la température baisse. Elle est bien moindre à la température ordinaire qu'à 500 degrés. À cette dernière température, l'aluminium peut retenir 5 % de cuivre en solution solide, tandis qu'à la température ordinaire, il ne peut en retenir que 2 %. Au cours d'un refroidissement lent de 500 degrés C à la température ordinaire, 3 % de cuivre sous la forme du composé CuAl^2 doivent donc être rejetés hors de la solution solide pour

(1) Voir Revue de Métallurgie, t. XXXIV, n° 3, mars 1937, p. 205, et n° 4, avril 1937, p. 253.

Metallurgical Reminiscences And Dialogue

Richard Bailey



Metallurgical Reminiscences And Dialogue:

The book delves into Metallurgical Reminiscences And Dialogue. Metallurgical Reminiscences And Dialogue is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Metallurgical Reminiscences And Dialogue, encompassing both the fundamentals and more intricate discussions.

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