

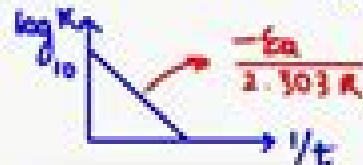
Temperature

$$A = \frac{K_2 + 10}{K_1}$$

$$\frac{k_2}{k_1} = \frac{K_2}{K_1}$$

Arrhenius equation

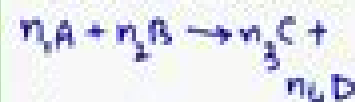
$$\log_{10} K = \log_{10} A - \frac{E_a}{2.303RT}$$



Collision theory

→ Reactants must have threshold Energy
→ Proper orientation

Rate of reaction



$$-\frac{1}{n_1} \frac{d[A]}{dt} = -\frac{1}{n_2} \frac{d[B]}{dt}$$

$$= \frac{1}{n_3} \frac{d[C]}{dt} = \frac{1}{n_4} \frac{d[D]}{dt}$$

Rate law $\rightarrow K[A]^x[B]^y$
order $\rightarrow x+y$

Nth order reaction

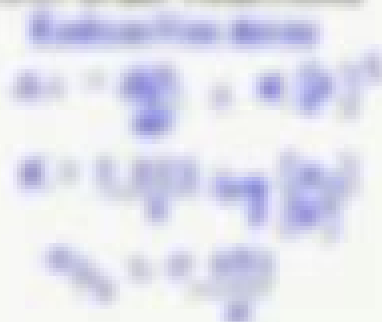
$$n \neq 1 \Rightarrow K_t = \frac{1}{(n-1)} \left[\frac{1}{[A]^{n-1}} - \frac{1}{[A]_0^{n-1}} \right]$$

$$t_{1/2} = \frac{1}{[A]_0^{n-1}}$$

Chemical kinetics

Molecularity

First order reactions



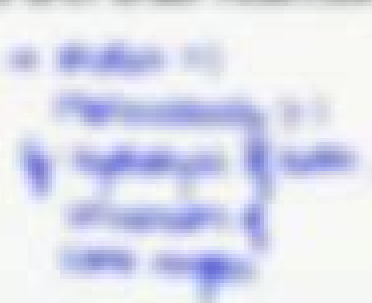
Zero order reaction



Elementary reactions

→ first order is unimolecular
→ second order is bimolecular
→ third order is termolecular

Overall first order reaction



Research In Chemical Kinetics

Daniel F McAuley



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