

General Chemistry/Health Sciences Worksheet on Chapter 10: The pH Scale

Some important equations to use are highlighted in bold below.

pH is a measure of a solution's hydronium ion concentration. It is defined as: $\text{pH} = -\log [\text{H}_3\text{O}^+]$

pOH is a measure of a solution's hydroxide ion concentration. It is defined as: $\text{pOH} = -\log [\text{OH}^-]$

The self-ionization of water: $2\text{H}_2\text{O}_{(l)} \rightleftharpoons \text{H}_3\text{O}^+_{(aq)} + \text{OH}^-_{(aq)}$

In pure water or neutral solutions, $[\text{H}_3\text{O}^+] = [\text{OH}^-] = 1 \times 10^{-7}$ and $\text{pH} = 7$

The concentrations of H_3O^+ and OH^- are related to one another as such: $[\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$

And therefore: $\text{pH} + \text{pOH} = 14$

These problems are taken from the textbook. They can be answered in the space provided or you can submit an image file (i.e., a picture) of your answers.

14.49 Calculate the pH and the pOH of each solution given the following:

✓ a. $[\text{H}_3\text{O}^+] = 1 \times 10^{-4} \text{ M}$

✓ b. $[\text{H}_3\text{O}^+] = 3 \times 10^{-9} \text{ M}$

✓ c. $[\text{OH}^-] = 1 \times 10^{-5} \text{ M}$

d. $[\text{OH}^-] = 2.5 \times 10^{-11} \text{ M}$

e. $[\text{H}_3\text{O}^+] = 6.7 \times 10^{-8} \text{ M}$

✓ f. $[\text{OH}^-] = 8.2 \times 10^{-4} \text{ M}$

14.50 Calculate the pH and the pOH of each solution given the following:

a. $[\text{H}_3\text{O}^+] = 1 \times 10^{-6} \text{ M}$

b. $[\text{H}_3\text{O}^+] = 5 \times 10^{-6} \text{ M}$

c. $[\text{OH}^-] = 1 \times 10^{-2} \text{ M}$

d. $[\text{OH}^-] = 8.0 \times 10^{-3} \text{ M}$

e. $[\text{H}_3\text{O}^+] = 4.7 \times 10^{-2} \text{ M}$

f. $[\text{OH}^-] = 3.9 \times 10^{-4} \text{ M}$

metabolic acidosis has a blood plasma pH of 6.92. What is the $[\text{H}_3\text{O}^+]$

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