



# Lower Manhattan New York Pocket Map

**Camilla Rothe**



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type of chemical reaction for this reaction we have a neutralization reaction balancing strategies here we have a  
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solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the  
gaseous hydrogen produced

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equation balancer balanced chemical equation  $2\text{hno}_3 \text{ mg c}_2\text{h}_3\text{o}_2 \text{ 2 2hc}_2\text{h}_3\text{o}_2 \text{ mg no}_3 \text{ 2}$

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reaction where one mole of solid magnesium mg and two moles of aqueous nitric acid hno 3 react to form one mole of  
aqueous magnesium nitrate  $\text{mg no}_3 \text{ 2}$  and one mole of dihydrogen h 2 gas show chemical

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molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of  
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