

1)

A) Ideal gas law is written as : $PV = nR'T$

Where, P = pressure, V = volume , T = temperature ,
 R' = universal gas constant and n = no. of moles of
the gas

It is also written as $PV = mRT$

Where m = mass of the gas and R = gas constant
for particular substance = R'/M

B) $T = 50^\circ\text{C}$

T (In kelvin) = $T + 273 = 323\text{ K} : \text{ANSWER}$

C) $P = 500\text{ kPa}$

$T = 323\text{ K}$

$m = 10\text{ kg}$

$R = 0.287\text{ kJ/kg K}$

Putting above values in ideal gas equation

$500 V = 10 \times 0.287 \times 323$

$V = 1.854\text{ m}^3 : \text{ANSWER}$

Rozners Constant

CO Houle



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