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water $n_a = 54 \text{ g} / 18 \text{ g mol}^{-1} = 3 \text{ mol}$ total number of moles $n_a + n_b = 0.5 + 3 = 3.5 \text{ mol}$ mole fraction of solute ethyl alcohol $x_b = n_b / n_a + n_b = 0.5 / 3.5 = 0.1429$

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40 g of H_2O molar mass of H_2O is 18 gm mole and molar mass of CH_3OH is 32 gm mole solution moles of $\text{CH}_3\text{OH} = 4.5 / 32 = 0.14$
mole moles of $\text{H}_2\text{O} = 40 / 18 = 2.2$ moles therefore according to the equation

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time the total pressure $p_a = x_a p_{\text{total}}$ mole fraction $x_a = \frac{\text{moles of } a}{\text{total moles}}$ mole fraction of substance
 a is often denoted by x_a and

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by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction
mole fraction is not dependent on the temperature

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