

# Kinetic Molecular Theory of Liquids

- 1. At normal atmospheric pressure, most substances are more dense in a liquid state than in a gaseous state.
- 2. Liquids are less compressible than gases
- 3. Liquids have the ability to diffuse, but it occurs slower than gases, because their particles are closer together.

Diffusion



As time goes on, the two liquids in this beaker diffuse into each other. The green liquid from coloring from the dye will eventually form a uniform solution with the water.

# Molecular Theory Of Gases And Liquids

**Joseph O. Hirschfelder**



## **Molecular Theory Of Gases And Liquids:**

**The Molecular Theory of Gases and Liquids** Joseph O. Hirschfelder, Charles F. Curtiss, R. Byron Bird, 1964-01-15 An essential cross disciplinary reference for molecular interactions Molecular Theory of Gases and Liquids offers a rigorous comprehensive treatment of molecular characteristics and behaviors in the gaseous and fluid states A unique cross disciplinary approach provides useful insight for students of chemistry chemical engineering fluid dynamics and a variety of related fields with thorough derivations and in depth explanations throughout Appropriate for graduate students and working scientists alike this book details advanced concepts without sacrificing depth of coverage or technical detail **Molecular**

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Theory of Gases and Liquids, 1967 *Molecular Theory of Nematic (and Other) Liquid Crystals* Paul van der Schoot, 2022-07-29 This book provides a didactic derivation of the main theories of thermotropic and lyotropic liquid crystals revealing the common molecular theoretic framework that underpins both theories This unified context will help young researchers in coming to grips with the basics of the simplest of liquid crystals being uniaxial nematic liquid crystals easing them into the intricacies of more complex forms of such materials irrespective of whether they are thermotropic or lyotropic The coverage provides a theoretical understanding of the phase behaviour that is what drives molecules and particles to spontaneously align themselves as well as an appreciation of the role of entropy energy and so on The focus here is on the main theories for the isotropic nematic transition being the Maier Saupe and the Onsager theories and how they are derived from a common description known as classical density functional theory DFT This book will be a valuable resource for senior

undergraduate and graduate students and experimentalists and engineers who feel intimidated by more formal or rigorous theoretical accounts and textbooks Exercises at the end of each chapter help the reader to apply the basic concepts also to other types of liquid crystal in particular the smectic liquid crystal

**Molecular Theory of Gases and Liquids. [By] J. O. Hirschfelder ... Charles F. Curtiss ... R. Byron Bird, etc** Joseph Oakland Hirschfelder, Robert Byron Bird, Charles Francis Curtiss, 1954 *A Kinetic Theory of Gases and Liquids* Richard Daniel Kleeman, 2013-09 This historic book may have numerous typos and missing text Purchasers can usually download a free scanned copy of the original book without typos from the publisher Not indexed Not illustrated 1920 edition Excerpt of gas from which the molecule comes and thus increases with decrease of pressure The difference between  $V_a$  and  $V_c$  therefore also increases with decrease of pressure which increases the value of the ratio  $T_j$   $i_j$  according to the foregoing equation This corresponds to a decrease of and thus the effect of slipping becomes the greater the lower the pressure The dynamical mechanism underlying the result that the viscosity of a gas is independent of its density may be GASEOUS VISCOSITY AND THE TEMPERATURE 127 illustrated by the following considerations A molecule transfers a certain amount of momentum to the gas at the end of each transfer distance on migrating at right angles to the motion of the gas in the direction of the decrease of motion If the concentration of the gas is halved the length of each transfer distance is doubled while the momentum transferred at the end of each transfer distance is also doubled since the velocity gradient of the gas remains the same Since a change in molecular concentration of a gas does not alter the molecular velocities the momentum transferred per second by a molecule moving between two parallel plates of material one of which is at rest while the other moves parallel to itself is in the latter case double that in the former But since the number of molecules per cubic cm available for momentum transference in the former case is half the number in the latter the total momentum transferred is in each case the same or the viscosity of the gas has not been altered by altering its density The quantity  $K$  is a function of the temperature This is shown by the calculated values of  $K$  contained in Table X for a number of gases at different temperatures These values are not corrected for

*Physical Chemistry for the Biosciences* Raymond Chang, 2005-02-11 This book is ideal for use in a one semester introductory course in physical chemistry for students of life sciences The author's aim is to emphasize the understanding of physical concepts rather than focus on precise mathematical development or on actual experimental details Subsequently only basic skills of differential and integral calculus are required for understanding the equations The end of chapter problems have both physiochemical and biological applications

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